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chapter - Chemical Kinetics - 01.

Admit
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i) Instantaneous Rate of Reaction.

- (a) It occurs within a short span of time.
- (b) It cannot be calculated for multistep ~~elementary~~ reaction.
- (c) It can be calculated for elementary reaction.

Average Rate of Reaction.

- (a) It occurs during a long interval of time.
- (b) It can be calculated for multistep reaction.
- (c) It cannot be calculated for elementary reaction.

ii) If the unit of rate constant is s^{-1} , the reaction is of first order.

iii) No, the molecularity can never be zero or a fractional number as it shows the number of reactants taking part in a reaction can never be zero.

iv)



Rate of rxn (M) = $k[R]^n$ → order of rxn. Reaction involves two reactants.

$x = k[R]^n$ — (1)

$8x = k[2R]^n$ — (2)

Dividing eq (2) by (1);

$$\frac{8x}{x} = \frac{k[2R]^n}{k[R]^n}$$

or, $8 = 2^n \cdot \frac{8M}{8M}$

Overall order

$= (a+b)$

$= (1+2) = 3$

or, $8 = 2^n$

or, $2^3 = 2^n$

$\therefore n = 3$

The order of reaction is 3.

(1) $x = k[A]^a[B]^b$

(2) $2x = k[2A]^a[B]^b$

(3) $8x = k[2A]^a[2B]^b$

(2) ÷ (1) →

$$\frac{2x}{x} = \frac{k[2A]^a[B]^b}{k[A]^a[B]^b}$$

$2^1 = 2^a \Rightarrow a = 1$

(3) ÷ (2) →

$2^b = 4 = 2^2 \Rightarrow b = 2$

v) The rate constant for the first order reaction is $60 s^{-1}$.

let M be the molar mass.
a be the initial concentration.

$k = 60/s$

final concentration = $\frac{M}{16} \times a$

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