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web a reaction and its experimental rate law are represented below $2 \text{ icl} + \text{h}_2 \text{g} \rightarrow 2 \text{ hcl} + \text{g}$ rate $k [\text{icl}]^2 [\text{h}_2]$ a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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web reaction rates google classroom $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ the reaction between Mg and HCl is represented by the equation above in a kinetics experiment a 0.080 g sample of Mg ribbon is added to an erlenmeyer flask containing 50 ml of 3.0 M HCl at 25 °C

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