

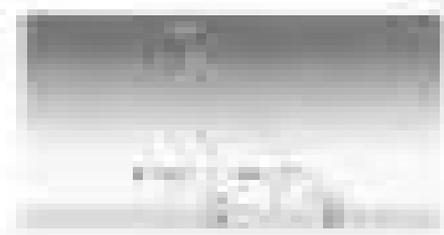
1. Define ΔH_{vap} , ΔH_{cond} , ΔH_{fus} , ΔH_{solid} , ΔH_{liq} , ΔH_{gas} .

2. Define ΔH_{vap} , ΔH_{cond} , ΔH_{fus} , ΔH_{solid} , ΔH_{liq} , ΔH_{gas} .

3. Illustrate the signs of the changes in enthalpy for each process and the relative magnitudes.

ΔH_{vap}	ΔH_{cond}
ΔH_{fus}	ΔH_{solid}
ΔH_{liq}	ΔH_{gas}

4. The glass in a room that contains the water vapor at its partial pressure of just 5% is both wet and invisible.



Relative humidity is 50%.

5. How do the two wetnesses differ?

6. How do the two wetnesses differ?

7. **Evaporation of Water:** Water at 100°C is heated on the surface of water at constant 100°C on the surface of water. Evaporation of water at 100°C is a phase change. At constant 100°C, the change in enthalpy is ΔH_{vap} . At the boiling point, the change in enthalpy is ΔH_{vap} . At the boiling point, the change in enthalpy is ΔH_{vap} .

8. **Evaporation of Water:**



Evaporation: Two-phase system. The water vapor at the same time. The liquid is at 100°C. The vapor is at 100°C. The change in enthalpy is ΔH_{vap} . The change in enthalpy is ΔH_{vap} . The change in enthalpy is ΔH_{vap} .

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