

Structure of H₂O
hydrogen bonding
Dimer of H₂O



unit one guide

Biological molecules
proteins
nucleic acids

H₂O provides cushion, adhesion, surface tension, high specific heat, high heat of vaporization, excellent solvent, stable chemistry, low compressibility

hydrogen bonds require lots of energy to change state of matter

CONDENSATION → when H₂O leaves it produces hydrogen and hydroxyl ions, releases pH → more hydrogen

OH⁻ → attracted to some
OH⁻ → attracted to other
OH⁻ → hydrogen bonds
OH⁻ → hydrogen bonds
OH⁻ → surface of the liquid
OH⁻ → surface of the liquid

C	H	N	O	P	S
Carbon	Hydrogen	Nitrogen	Oxygen	Phosphorus	Sulfur

more hydrogen
C → build biological molecules, cell formation
N → build proteins, nucleic acids
P → build nucleic acids + lipids
→ make up 98% of organisms

Carbohydrates ↓ C, H, O	Lipids ↓ C, H, O, P	Nucleic acids ↓ C, H, N, O, P	Proteins ↓ C, H, N, O, S
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Condensation synthesis → the removal of an H₂O molecule to bond monomers. Hydrolysis → the addition of a water molecule to cleave a bond between monomers (breaks protein)

Nucleic acid structure:

• the carbon sugar (deoxyribose/ribose), phosphate, nitrogenous base (A, T, C, G), exact shape of strand



Starch vs Cellulose

Starch → α glucose
Cellulose → β glucose

Proteins:

Specific order of amino acids → primary structure

→ 1 group interactions determine structure + function

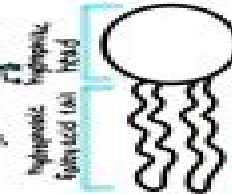


PROTEINS

• amino acids bond via a peptide bond (covalent) through dehydration synthesis
• 1 group change can cause a hydrophobic protein to lose the place of a hydrophilic protein

membranes

• made up of phospholipids
• create selectively permeable barrier
• contain phospholipids, cholesterol, and proteins



LIPIDS

Phospholipid macromolecules
• differences in saturation determine structure and function
• Cells, sterols, phospholipids
• Saturated → no double bonds
• Unsaturated → has double bonds
• Lipids: triglycerides, cholesterol, most of the membrane proteins, 4 times carbon rings, cholesterol, maintaining flexibility, embedded into up cell membranes, polar region due to phosphorus and fatty acid



Amino Acids:

• have directionality
• linear chains, connected by covalent bonds
• amino end and carboxyl end

DEHYDRATION SYNTHESIS



Chaperones:

• protein molecules that assist protein folding

NUCLEIC ACIDS

• add new nucleotides to the 3' end
• base nitrogenous bases: purine + pyrimidine
• pyrimidine is single ringed, purine is double
• antiparallel double helix
• 5' on phosphate end, 3' on hydroxyl end
• both base and back have nitrogenous bases (purines) + sugar-phosphate backbone

ANTIPARALLEL



ALDOSE VS KETOSE

• contains aldehyde functional group



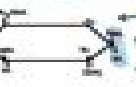
KETOSE

• contains ketone functional group



ALPHA VS BETA

• all monomers are in the same orientation



BETA

• Every glucose monomer is upside down from its neighbor

NH₂ = Amino
COOH = Carboxyl

PROTEIN STRUCTURE

• Primary: sequence of amino acids, specific order determine protein shape
• Secondary: how folding of chain into alpha helix and beta sheet occurs
• Tertiary: 3-D shape, hydrogen bonds, disulfide link, ionic bonds, hydrophobic interactions
• Quaternary: interactions between multiple polypeptide chains

+ FUNCTION

• Enzymes → speeds up reaction rates in the cell
• Cell structure → part of the cellular membrane
• Communication between cells

DENATURATION:

• the protein starts to lose its shape/structure when (bonds break)
• Temperature change
• pH of surrounding area
• Salt concentration

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