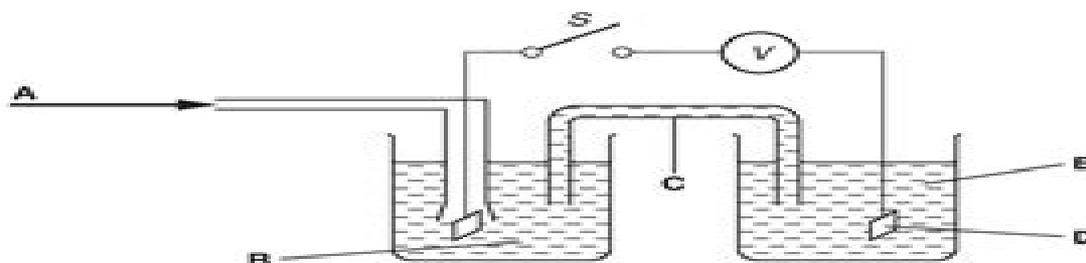


## ELECTROCHEMISTRY WS 1

- 1 Chlorine gas and iron(II) ions react together in aqueous solution as follows.



- (a) The following diagram shows the apparatus needed to measure the  $E_{\text{cell}}^{\circ}$  for the above reaction.



- (i) In the spaces below, identify what the five letters A – E in the above diagram represent.

A ..... chlorine at  $P_{\text{Cl}_2} = 1 \text{ atm}$  and  $298 \text{ K}$   
 B .....  $1 \text{ mol dm}^{-3} \text{ NaCl}$   
 C ..... salt bridge  
 D ..... Platinum  
 E .....  $1 \text{ mol dm}^{-3} \text{ Fe}^{2+} \text{ aq}$  and  $1 \text{ mol dm}^{-3} \text{ Fe}^{3+} \text{ aq}$

- (ii) Use the *Data Booklet* to calculate the  $E_{\text{cell}}^{\circ}$  for this reaction, and hence decide which direction (left to right, or right to left) electrons would flow through the voltmeter V when switch S is closed.



[7]

# Electrochemistry Answers

**Ponnadurai Ramasami**



## **Electrochemistry Answers:**

*Solutions Manual to Accompany Elements of Physical Chemistry* David Smith, 2013-05-30 The Solutions Manual to accompany Elements of Physical Chemistry 6th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions

**Electrochemistry** Christine Lefrou, Pierre Fabry, Jean-Claude Poignet, 2012-05-24 This textbook offers original and new approaches to the teaching of electrochemical concepts principles and applications Throughout the text the authors provide a balanced coverage of the thermodynamic and kinetic processes at the heart of electrochemical systems The first half of the book outlines fundamental concepts appropriate to undergraduate students and the second half gives an in depth account of electrochemical systems suitable for experienced scientists and course lecturers Concepts are clearly explained and mathematical treatments are kept to a minimum or reported in appendices This book features Questions and answers for self assessment Basic and advanced level numerical descriptions Illustrated electrochemistry applications This book is accessible to both novice and experienced electrochemists and supports a deep understanding of the fundamental principles and laws of electrochemistry

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**Electrochemistry** John O'M. Bockris, 1973

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In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Electrochemistry Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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