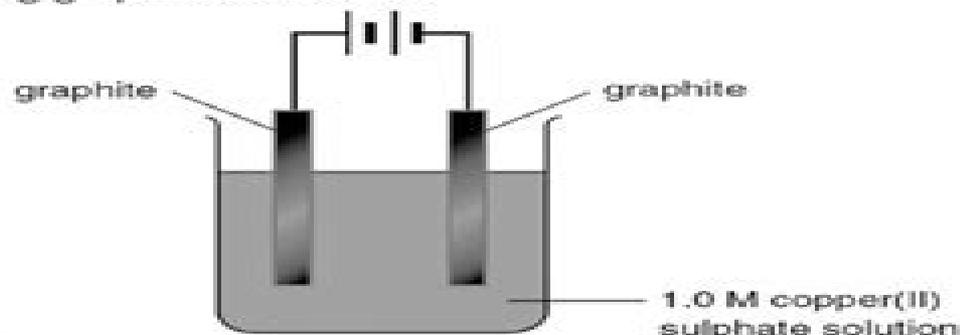


### Quiz (Electrolysis – Further Exercise)

1. The diagram on the right shows the electrolysis of a 1.0 M copper(II) sulphate solution using graphite electrodes.



- (a) Write the half equations for reactions taking place at the anode and the cathode respectively.
- (b) When electrolysis is completed, a few drops of universal indicator are added to the solution. State the colour of the resultant solution.
- (c) The electrolyte is replaced by 5.0 M copper(II) sulphate solution. State the product formed at each electrode. Explain your answer.
2. The diagram below shows a set-up for the electrolysis of concentrated sodium bromide solution. After passing electricity for some time, gas bubbles were observed at electrode *P*.
- (a) Which of the electrodes, *P* or *Q*, is the cathode?
- (b) Name the gas produced at electrode *P* during electrolysis, and write a half equation for the reaction that occurred.
- (c) What ion(s) would migrate towards electrode *Q* during electrolysis?
- (d) A yellowish brown colour appeared around electrode *Q* during electrolysis. Write a half equation for the reaction that occurred.
- (e) The yellowish brown colour gradually extended from electrode *Q* to the bottom of the U-tube. Explain briefly.
- (f) Suggest TWO ways to speed up the electrolysis, using the same set-up and chemicals.

# Electrolysis Questions And Answers

**Amar Nath Bhutani**



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