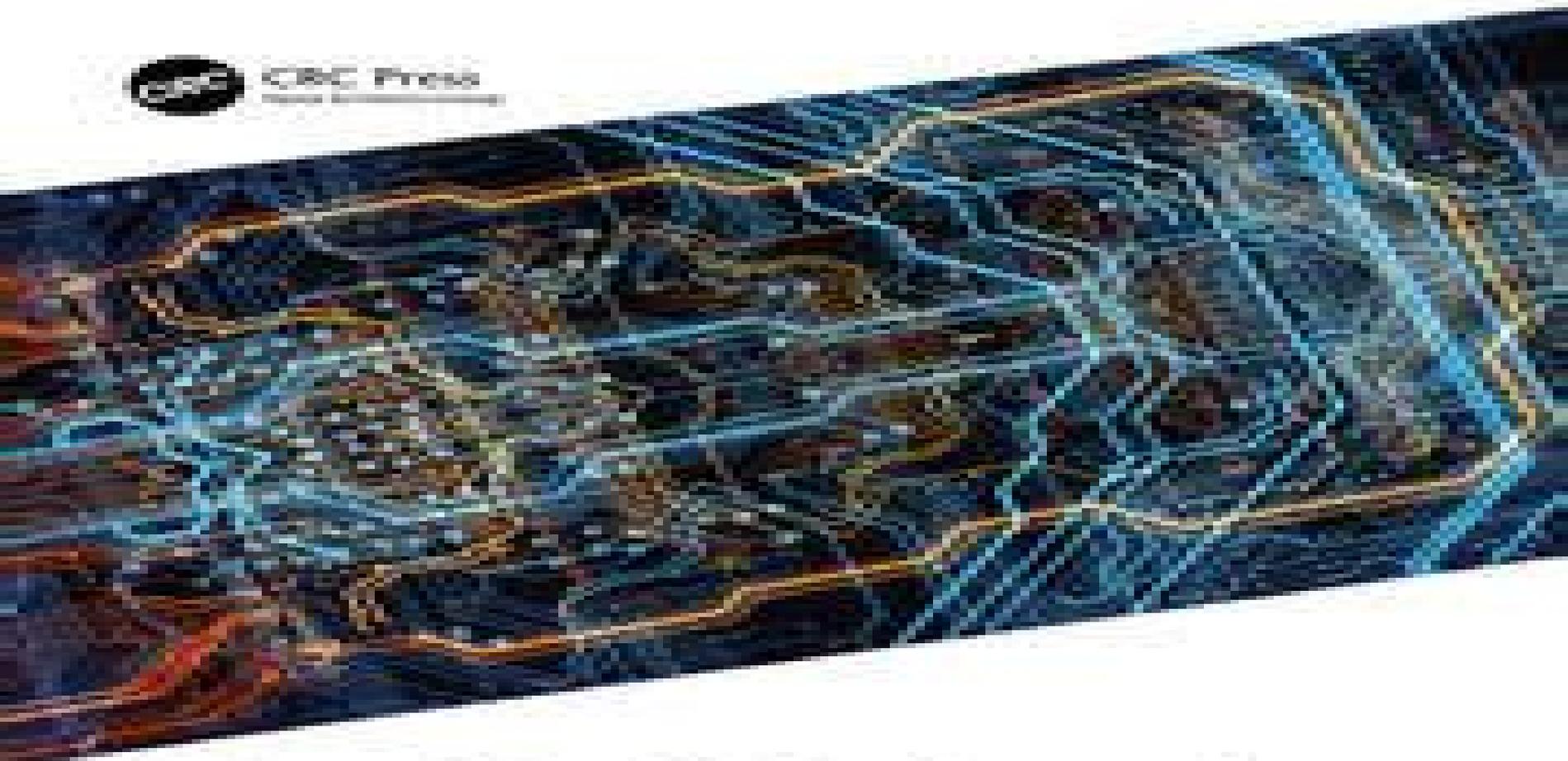




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# Electrical Engineering Fundamentals

S. Bobby Rauf

# Electrical Engineering Fundamentals

**Giorgio Rizzoni**



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aspects of electrical science including Ohm and Kirchoff's laws P N junctions semiconductors circuit diagrams magnetic fields electrochemistry and devices such as DC motors This text is useful for students of electrical chemical materials and mechanical engineering

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**Electrical Engineering Fundamentals** George Francis Corcoran, Edwin Bernard Kurtz, 1941

*ISE Fundamentals of Electrical Engineering* Giorgio Rizzoni, 2021-01-12

Fundamentals of Electrical Engineering Charles A. Gross, Thaddeus A. Roppel, 2012-02-15. Real world engineering problems are rarely, if ever, neatly divided into mechanical, electrical, chemical, civil, and other categories. Engineers from all disciplines eventually encounter computer and electronic controls and instrumentation, which require at least a basic knowledge of electrical and other engineering specialties, as well as associated economics and environmental, political, and social issues. Co-authored by Charles Gross, one of the most well-known and respected professors in the field of electric machines and power engineering, and his world-renowned colleague Thad Roppel, *Fundamentals of Electrical Engineering* provides an overview of the profession for engineering professionals and students whose specialization lies in areas other than electrical. For instance, civil engineers must contend with commercial electrical service and lighting design issues; mechanical engineers have to deal with motors in HVAC applications; and chemical engineers are forced to handle problems involving process control. Simple and easy to use, yet more than sufficient in rigor and coverage of fundamental concepts, this resource teaches EE fundamentals but omits the typical analytical methods that hold little relevance for the audience. The authors provide many examples to illustrate concepts, as well as homework problems to help readers understand and apply presented material. In many cases, courses for non-electrical engineers or non-EEs have presented watered-down, classical EE material, resulting in unpopular courses that students hate and senior faculty members understandably avoid teaching. To remedy this situation and create more well-rounded practitioners, the authors focus on the true EE needs of non-EEs, as determined through their own teaching experience, as well as significant input from non-EE faculty. The book provides several important contemporary interdisciplinary examples to support this approach. The result is a full-color, modern narrative that bridges the various EE and non-EE curricula and serves as a truly relevant course that students and faculty can both enjoy.

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web google classroom for the following redox equation  $\text{MnO}_4^{-} + \text{aq} + \text{SO}_2 + \text{Mn}^{2+} + \text{aq} + \text{H}^{+} + \text{aq}$  what will be the stoichiometric coefficient of  $\text{H}^{+}$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{H}^{+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+} + \text{H}_2\text{O}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} \rightarrow 2\text{H}_2\text{O} + \text{H}_2\text{O}_2 + 2\text{OH}^{-}$  make electrons equal convert to basic solution  $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-}$  multiplied by a factor of 2  $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 + \text{H}_2\text{O}_2 + 2\text{OH}^{-}$  3 the final answer

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h<sub>2</sub>o

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web bonus problem here is a balanced net ionic equation  $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$  the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}^{2+}$   $\text{Fe}^{2+}$   $\text{Cr}^{3+}$   $\text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^-$   $14\text{H}^+$   $\text{Cr}^{2+}$   $2\text{Cr}^{3+}$   $7\text{H}_2\text{O}$   $\text{Fe}^{2+}$   $\text{Fe}^{3+}$   $\text{e}^-$  2 equalize the electrons  $6\text{e}^-$   $14\text{H}^+$   $\text{Cr}^{2+}$   $2\text{Cr}^{3+}$   $7\text{H}_2\text{O}$   $6\text{Fe}^{2+}$   $6\text{Fe}^{3+}$   $6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}^{2+}$   $\text{NO}_3^-$   $\text{SO}_4^{2-}$   $\text{MnO}_2$   $\text{Cr}^{3+}$   $\text{NO}_3^-$   $\text{SO}_2$   $\text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2$   $\text{Cr}^{2+}$   $\text{TeO}_3$   $\text{N}_2\text{O}_4$   $\text{O}_2$   $\text{Cr}^{3+}$   $\text{Te}$  10  $\text{ReO}_4^-$  11  $\text{PbO}_2$   $\text{IO}_3^-$   $\text{IO}_2$   $\text{Pb}^{2+}$  12  $\text{AsO}_3^{3-}$   $\text{IO}_3^-$   $\text{Re}$   $\text{IO}_3^-$   $\text{H}_2\text{AsO}_4^-$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6$   $2\text{H}_2\text{NO}_2$   $\text{C}_6\text{H}_6\text{O}_6$   $2\text{H}_2\text{O}_2$   $2\text{NO}$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  b  $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  c  $\text{H}_2\text{O}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  d  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{H}_2\text{O}_2$

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