

# Half Modern Chemistry Review

## CHAPTER 9: STOICHIOMETRY

The following pages contain the best (but not all) of the information for the chapter 9 test.

Focus on this content, but make sure to review class notes, activities, worksheets, questions, etc.

If you study this document and NOTHING else, you should at least be able to PASS the test.

\*\*\*\* Test items will be recall, recognition, and/or application of this content. \*\*\*\*

### OBJECTIVES

- Calculate molar weight for molecular-formula compounds. (C 1)
- Calculate molar weight for ionic compounds. (C 1)
- 9.1. Determine moles of reactants from balanced chemical equations. (T & P)
- 9.2. Perform stoichiometry calculations with: mol to mol, mol to gram, gram to mol, and gram to gram. (P & P)
- 9.3. Calculate theoretical yield and determine the limiting reagent, excess reagent, and amount of product produced. (P & P)

### 9.1: INTRODUCTION TO STOICHIOMETRY

#### a. Introductory

- composition stoichiometry – calculations involving the mass relationships of elements in compounds.
- reaction stoichiometry – calculations involving the mass relationships between reactants and products in a chemical reaction.
- molar ratio – a conversion factor that relates the amount, in moles, of any two substances involved in a chemical reaction.

#### b. Chapter Highlights

- Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.
- Stoichiometry is based on the law of conservation of mass.
- Reacting mass substances is specified along with expressing the amount of each substance in moles.
- A molar ratio in the chemical reaction that relates the amount in moles of any two substances in a chemical reaction. This molar ratio is derived from the balanced equation.
  - Example:  $2\text{FeCl}_3 + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_4 + 3\text{Cl}_2$  (Note: the ratio of iron to chlorine moles is 2 to 3)
- Amount of a substance is expressed in moles, and mass of a substance is expressed using mass units such as grams, kilograms, or milligrams.
- Mass and amount of substances are quantities, whereas moles and grams are units.
- A balanced chemical equation is necessary to solve any stoichiometric problem.

### 9.2: BASIC STOICHIOMETRIC CALCULATIONS

#### a. Chapter Highlights

- In an ideal stoichiometric calculation, the mass of the amount of any reactant or product can be calculated if the balanced chemical equation and the mass or amount of any other reactant or product is known.

### 9.3: LIMITED REACTANTS AND PERCENTAGE YIELD

# Modern Chemistry Chapter 9 Section 1 Review Answers

**CL Gary**



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## Enjoying the Tune of Expression: An Psychological Symphony within **Modern Chemistry Chapter 9 Section 1 Review Answers**

In some sort of taken by displays and the ceaseless chatter of fast connection, the melodic splendor and mental symphony created by the published word often diminish into the back ground, eclipsed by the persistent sound and interruptions that permeate our lives. But, nestled within the pages of **Modern Chemistry Chapter 9 Section 1 Review Answers** an enchanting fictional prize full of fresh thoughts, lies an immersive symphony waiting to be embraced. Crafted by an elegant composer of language, this charming masterpiece conducts readers on a psychological journey, well unraveling the hidden songs and profound affect resonating within each carefully crafted phrase. Within the depths with this poignant assessment, we shall examine the book is key harmonies, analyze their enthralling publishing fashion, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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