

Name _____

Molecular Geometries Lab Report Sheet

Formula	Lewis Structure	No of electron domains/lig	Electron domain geometry	VSEPR Molecular geometry	Hybridization of central atom	Polar ?
CO ₂	<p>Valence 24</p> $\text{O}=\text{C}=\text{O}$ <p>2 charge clouds</p>	0	Linear 180°	Linear 	sp	NP
BF ₃	<p>Valence 32</p> $\text{F}=\text{B}-\text{F}$	0	Trigonal planar 120°		sp ²	NP
SnCl ₂	<p>Valence 24</p> $\text{Cl}=\text{Sn}-\text{Cl}$	3	Bent 120°		sp ²	P
BF ₄ ⁻	<p>Valence 40</p> $\left[\text{F}-\text{B}-\text{F} \right]^-$	0	Tetrahedral 109.5°		sp ³	P
AsH ₃	<p>Valence 24</p> $\text{H}-\text{As}-\text{H}$	4	Trigonal pyramidal 109.5°		sp ³	P
H ₂ S	<p>Valence 14</p> $\text{H}-\text{S}-\text{H}$	2	Bent 109.5°		sp ³	NP
NH ₂ ⁻	<p>Valence 10</p> $\left[\text{H}-\text{N}-\text{H} \right]^-$	2	Bent 109.5°		sp ³	P
PCl ₄	<p>Valence 40</p> $\text{F}-\text{P}-\text{Cl}$	0	Trigonal bipyre 120 + 90		sp ³ d	P

Molecular Shapes Lab Activity Answers

JR Anderson



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