

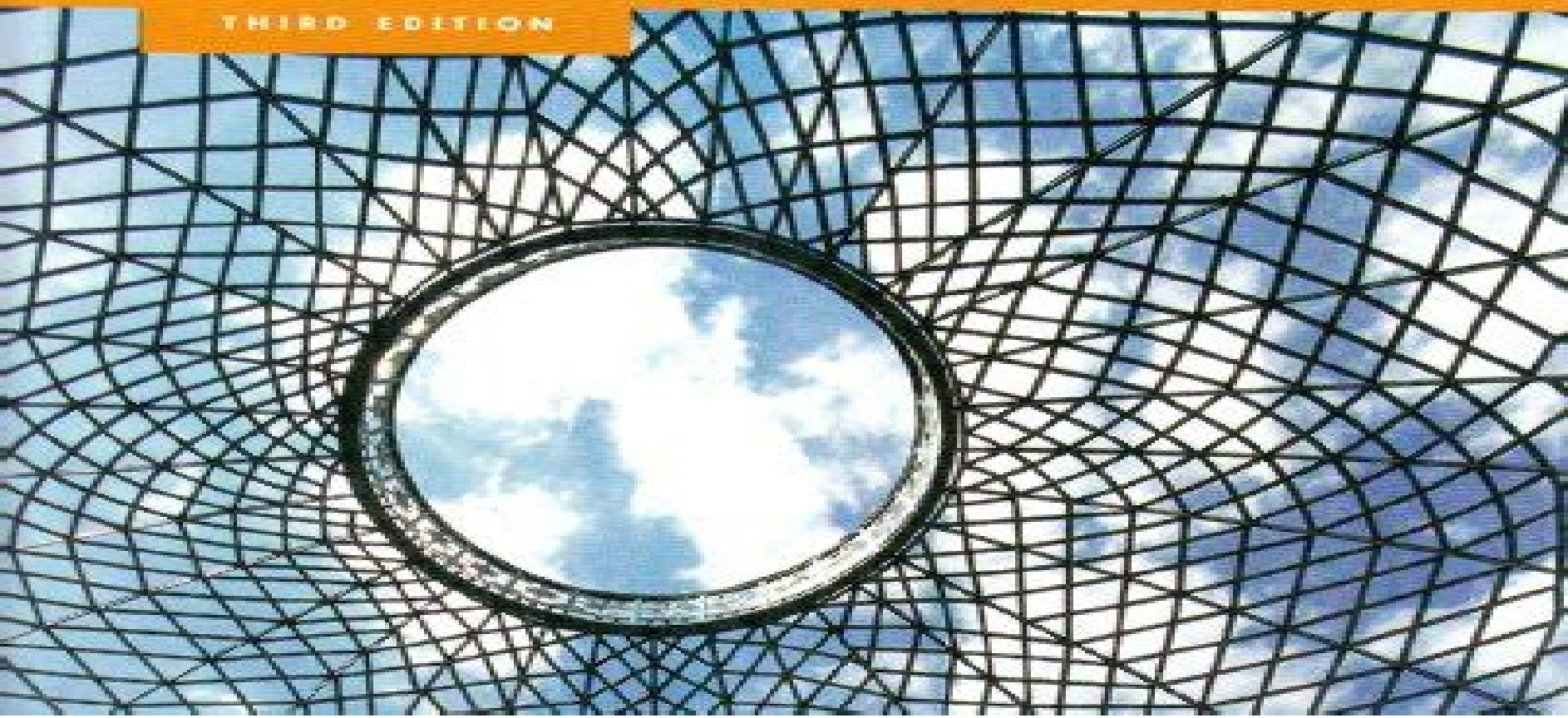
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Camilla Rothe



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web solution average atomic mass of li 6.941 amu x 6.0151 amu 1 x 7.0160 amu 6.941 1.0009 x 7.0160 1.0009 x 0.075 x 0.075 x 0.075 corresponds to a natural abundance of 6li of 7.5 percent the natural abundance of 7li is x 0.925 or 92.5 percent 3.7 the unit factor required is 6.022

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web a molar mass li₂co₃ 2 6.941 g 12.01 g 3 16.00 g 73.89 g b molar mass cs₂ 12.01 g 2 32.07 g 76.15 g c molar mass chcl₃ 12.01 g 1 008 g 3 35.45 g 119.37 g d molar mass c₆h₈o₆ 6 12.01 g 8 1.008 g 6 16.00 g 176.12 g e molar mass kno₃ 39.10 g 14.01 g 3 16.00 g 101.11 g

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6.4 mass relationships and chemical equations - Oct 03 2023

web it is a small step from mole mass calculations to mass mass calculations if we start with a known mass of one substance in a chemical reaction instead of a known number of moles we can calculate the corresponding masses of other substances in the reaction

answer key week 6 lesson 6.2 mass relationships in - Jun 18 2022

web topic mass relationships in chemical reactions materials textbook iupac periodic table scientific calculator pen and paper exercise a instruction after watching the two video files videos 24 and 25 solve the problems by yourself without looking into the video anymore items 1 2

3.4 mass relationships in chemical equations - Sep 02 2023

web jul 4 2022 to calculate the mass of titanium metal that you can obtain multiply the number of moles of titanium by the

molar mass of titanium 47.867 g/mol moles Ti mass Ti times molar mass Ti $\frac{47.867 \text{ g Ti}}{1 \text{ mol Ti}}$
197 g Ti notag

chapter 3 mass relationships in chemical reactions - Nov 23 2022

web 41 42 chemistry ch 3 mass relationships in chemical reactions molecular mass of NO_2 atomic mass of N 2 atomic mass of O 14.01 amu 2 16.00 amu 46.01 amu atomic mass of N_2O_5 2 atomic mass of N 5 atomic mass of O 2 14.01 amu 5 16.00 amu 108.02 amu example 3 2 exercise 3 3