

## Chapter 6 Covalent Bonds

**Section 6.2 Covalent Bonding****(pages 166–188)**

This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules.

**Reading Strategy (page 166)**

**Relating Text and Visuals** As you read the section, look closely at Figure 9. Complete the table by describing each type of model shown. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Molecular Models	
Model	Description
Electron-dot diagrams	Dots represent valence electrons.
Structural formula	
Space-filling	
Electron cloud	

**Covalent Bonds (pages 166–187)**

- Define a covalent bond. \_\_\_\_\_  
\_\_\_\_\_
- A molecule is a \_\_\_\_\_ group of atoms that are joined together by one or more covalent bonds. Circle the correct answer.  
negative      neutral      positive
- Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. \_\_\_\_\_
- Circle the correct answer. Nitrogen has five valence electrons. How many pairs of electrons must two nitrogen atoms share in order for each atom to have eight valence electrons?
  - one
  - two
  - three

# Chapter 6 Chemical Bonding Section 2 Covalent Answer Key

**Y Pai**



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