

CHAPTER 4 REVIEW

Chemical Equations and Reactions

Short Review

Write a balanced chemical equation for the reaction described.

1. _____
 - a. balanced chemical equation represents all the following except:
 - i. experimentally established facts
 - ii. observations by which reactions can be predicted
 - iii. identities of reactants and products in a chemical reaction
 - iv. relative quantities of reactants and products in a chemical reaction
 - b. According to the law of conservation of mass, the total mass of the
 - i. reacting substances is
 - i. always more than the total mass of the products
 - ii. always less than total mass of the products
 - iii. sometimes more and sometimes less than the total mass of the products
 - iv. always equal to total mass of the products
2. For the reaction described below, balanced equations will occur. For each reaction that will occur, identify the reaction type and complete the chemical equation, naming the products. Indicate and describe the final product. (Observability rules can be found in Figure 4.3 of the text.)
 - a. $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow$

 - b. $\text{Al} + \text{H}_2\text{SO}_4 \rightarrow$

 - c. $\text{Li}_2\text{O} + \text{H}_2\text{SO}_4 \rightarrow$

 - d. $\text{K}_2\text{SO}_4 + \text{CaCl}_2 \rightarrow$

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