

## CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron (II) sulfide, the roman numeral tells us (2)
  - How many atoms of Fe are in one formula unit.
  - How many sulfide ions can be attached to the iron atom.
  - The charge on each Fe ion.
  - The total positive charge of the formula unit.
- The result of changing a subscript in a correctly written chemical formula is (1)
  - Change the number of ions represented in the formula.
  - Change the charges on the other ions in the compound.
  - Change the formula so that it no longer represents the compound.
  - Have no effect on the formula.
- Derive a generalization for when an acid name will end in the suffix -ic or -ous. If ending in -ic, the anion ended in -ate. If ending in -ous, the anion ended in -ite.
- Derive a generalization for when an acid name will begin with the prefix hydro-. If prefix is hydro-, the anion ended in -ite. This is almost always a monatomic ion.
- Fill in the blanks in the table below.

Compound Name	Formula
Acetic acid	$\text{CH}_3\text{CO}_2\text{H}$
Aluminum sulfate	$\text{Al}_2(\text{SO}_4)_3$
Lead II carbonate	$\text{PbCO}_3$
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Sulfuric acid	$\text{H}_2\text{SO}_4$
Nitrogen pentoxide	$\text{N}_2\text{O}_5$
Copper II carbonate	$\text{CuCO}_3$
Sodium sulfate	$\text{Na}_2\text{SO}_4$
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Mercury sulfate	$\text{HgSO}_4$
Nitrous acid	$\text{HNO}_2$
Magnesium perchlorate	$\text{Mg}(\text{ClO}_4)_2$
Iron II nitrate	$\text{Fe}(\text{NO}_3)_2$
Iron III nitrate	$\text{Fe}(\text{NO}_3)_3$
Carbon II oxide	$\text{CO}$
Sulfur dioxide	$\text{SO}_2$
Sulfur pentachloride	$\text{SCl}_5$
Tetraphosphorus trisulfide	$\text{P}_4\text{S}_3$
Carbonic dioxide	$\text{CO}_2$

- Name the following compound and identify its molecular ions or acid.
  - $\text{Ca}(\text{NO}_3)_2$ , calcium (II) nitrate - II
  - $\text{Fe}(\text{NO}_3)_3$ , Iron (III) nitrate - III
  - $\text{Fe}(\text{OH})_2$ , Iron (II) hydroxide - II
  - $\text{N}_2\text{O}_5$ , Nitrogen pentoxide - III

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**M Walker**



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