

Section 12.1: Equilibrium

Section 12.1: Equilibrium

Section 12.1: Equilibrium: Answer the following questions in the space provided.

1. _____
 - a. _____
 - (i) _____
 - (ii) _____
 - (iii) _____
 - (iv) _____
 - b. _____
 - (i) _____
 - (ii) _____
 - (iii) _____
 - (iv) _____
2. _____
 - a. _____
 - (i) _____
 - (ii) _____
 - (iii) _____
 - (iv) _____
 - b. _____
 - (i) _____
 - (ii) _____
 - (iii) _____
 - (iv) _____
3. _____
 - a. _____
 - (i) _____
 - (ii) _____
 - (iii) _____
 - (iv) _____
 - b. _____
 - (i) _____
 - (ii) _____
 - (iii) _____
 - (iv) _____

4. Given the reaction represented by the following balanced equation:



a. Balance the equation:



b. Write the expression for K_c for H_2O .

c. In a closed 10 L flask there are 0.10 mol H_2O , 0.05 mol CO_2 , and 0.05 mol H_2 .

d. The equilibrium constant of H_2O is 0.10 at 1000 K. What is the concentration of H_2 at 1000 K?

e. The equilibrium constant of H_2O is 0.10 at 1000 K. What is the concentration of CO_2 at 1000 K?

Chapter 9 Stoichiometry Section 1 Answers

S Marginson



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