

WORKSHEET
CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1
 Molecular formula of hydrocarbon or a mixture of hydrocarbons

ANSWERS KEY

QUESTION 1. Calculate the empirical formula of hydrocarbon and its molecular formula.

1. A hydrocarbon of 0.100 mol is completely burnt by 0.140 mol oxygen which volume measured at STP. **Find the molecular formula of hydrocarbon if pressure is not atmospheric pressure?**
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) compare the volume of hydrocarbon and oxygen as they are burnt at STP
 $0.100 \text{ mol} : 0.140 \text{ mol} = 1 : 1.4$
 - (c) deduce the molecular formula of hydrocarbon from the above equation
 $C_1H_{1.4}$

2. A hydrocarbon hydrocarbon contains 83.26% carbon and contains the carbon dioxide gas completely converted. The hydrocarbon reacts with oxygen to form carbon dioxide and water vapor. These gases cooling of constant pressure are measured by 1.00 g of hydrocarbon.
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) calculate the molecular weight of hydrocarbon from the above equation with constant pressure at STP
 - (c) deduce the molecular formula of hydrocarbon from the above equation at STP

3. Hydrocarbon is burnt in 1.00 g of oxygen gas. The hydrocarbon is completely burnt and the hydrocarbon is $CO_2 + H_2O$
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) how many moles of hydrocarbon are produced by the combustion of 1.00 g of hydrocarbon
 0.01 mol
 - (c) how many moles of hydrocarbon are produced from 1.00 g of hydrocarbon
 0.01 mol

QUESTION 2. Calculate the empirical formula of hydrocarbon.

1. The total amount of carbon dioxide and water produced by burning 0.100 mol of hydrocarbon containing 83.26% carbon and 16.74% hydrogen is 0.100 mol. Find the empirical formula of hydrocarbon.
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) compare the moles of hydrocarbon and oxygen

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