

- A.  $n = 4, l = 3, m_l = -2, m_s = +1/2$
- B.  $n = 3, l = 2, m_l = -3, m_s = -1/2$
- C.  $n = 3, l = 0, m_l = 0, m_s = +1/2$
- D.  $n = 4, l = 1, m_l = 1, m_s = -1/2$
- E.  $n = 2, l = 0, m_l = 0, m_s = +1/2$

39. The orbital diagram for a ground-state oxygen atom is

- A.  $1s \uparrow\downarrow \quad 2s \uparrow\downarrow \quad 2p \uparrow\downarrow \uparrow \_$
- B.  $1s \uparrow\downarrow \quad 2s \uparrow \quad 2p \uparrow \uparrow \uparrow$
- C.  $1s \uparrow\downarrow \quad 2s \uparrow\downarrow \quad 2p \uparrow \uparrow \uparrow$
- D.  $1s \uparrow\downarrow \quad 2s \uparrow\downarrow \quad 2p \uparrow\downarrow \uparrow \uparrow$
- E.  $1s \uparrow\downarrow \quad 2s \uparrow\downarrow \quad 2p \uparrow\downarrow \uparrow\downarrow \uparrow$

40. Which of the following ground-state electron configurations are correct for the atom or ion shown?

- A.  $\text{Cl}^-: [\text{Ne}]3s^23p^4$
- B.  $\text{Al}^{3+}: [\text{Ne}]$
- C.  $\text{As}: [\text{Ar}]4s^24p^3$
- D.  $\text{Ti}: [\text{Ar}]4s^24d^2$

41. In what row and group of the periodic table would you find the element with the electron configuration  $[\text{Kr}]5s^24d^{10}5p^2$ ?

- A. row 4, group 5A
- B. row 4, group 4A
- C. row 5, group 5A
- D. row 5, group 4A

42. Which one of the following pairs are isoelectronic?

- A.  $\text{Mn}^{2+}$  and Ar
- B.  $\text{K}^+$  and  $\text{Cl}^-$
- C.  $\text{Zn}^{2+}$  and  $\text{Cu}^{2+}$
- D.  $\text{Na}^+$  and  $\text{K}^+$
- E.  $\text{Cl}^-$  and S

43. Which of the atoms listed below has the largest radius?

- A. Cl
- B. I

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**Siddappa N.Byrareddy**



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