

## Chemistry Notes for class 12 Chapter 2 Solutions

Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution.

### Solute and Solvent

In a binary solution, solvent is the component which is present in large quantity while the other component is known as solute.

### Classification of Solutions

(A) Following types of solutions are seen on the basis of physical state of solute and solvent.

S.No.	Solute	Solvent	Examples
<b>Solid solutions</b>			
1	Solid	Solid	Alloys
2	Liquid	Solid	Hydrated salts, Amalgam of Hg with Na
3	Gas	Solid	Dissolved gases in mineral
<b>Liquid solutions</b>			
4	Solid	Liquid	Salt/sugar solution in water
5	Liquid	Liquid	Alcohol in water
6	Gas	Liquid	Aerated drinks, $O_2$ in water
<b>Gaseous solutions</b>			
7	Solid	Gas	Iodine vapour in air
8	Liquid	Gas	Water vapour in air
9	Gas	Gas	Air ( $O_2 + N_2$ )

[if water is used as a solvent, the solution is called aqueous solution and if not, the solution is called non-aqueous solution.]

(B) Depending upon the amount of solute dissolved in a solvent we have the following types of solutions:

(i) **Unsaturated solution** A solution in which more solute can be dissolved without raising temperature is called an unsaturated solution.

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### **Table of Contents Chemistry Chapter 11 Assessment Answers**

1. Understanding the eBook Chemistry Chapter 11 Assessment Answers
  - The Rise of Digital Reading Chemistry Chapter 11 Assessment Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 11 Assessment Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Chemistry Chapter 11 Assessment Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 11 Assessment Answers
  - Personalized Recommendations
  - Chemistry Chapter 11 Assessment Answers User Reviews and Ratings

- Chemistry Chapter 11 Assessment Answers and Bestseller Lists
- 5. Accessing Chemistry Chapter 11 Assessment Answers Free and Paid eBooks
  - Chemistry Chapter 11 Assessment Answers Public Domain eBooks
  - Chemistry Chapter 11 Assessment Answers eBook Subscription Services
  - Chemistry Chapter 11 Assessment Answers Budget-Friendly Options
- 6. Navigating Chemistry Chapter 11 Assessment Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 11 Assessment Answers Compatibility with Devices
  - Chemistry Chapter 11 Assessment Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 11 Assessment Answers
  - Highlighting and Note-Taking Chemistry Chapter 11 Assessment Answers
  - Interactive Elements Chemistry Chapter 11 Assessment Answers
- 8. Staying Engaged with Chemistry Chapter 11 Assessment Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 11 Assessment Answers
- 9. Balancing eBooks and Physical Books Chemistry Chapter 11 Assessment Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 11 Assessment Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 11 Assessment Answers
  - Setting Reading Goals Chemistry Chapter 11 Assessment Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 11 Assessment Answers
  - Fact-Checking eBook Content of Chemistry Chapter 11 Assessment Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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