

Chapter 8

Organic Chemistry – Some Basic Principles and Techniques

Question 8.1:

What are hybridisation states of each carbon atom in the following compounds?

$\text{CH}_2 = \text{C} = \text{O}$, $\text{CH}_3\text{CH} = \text{CH}_2$, $(\text{CH}_3)_2\text{CO}$, $\text{CH}_2 = \text{CHCN}$, C_6H_6

Answer 8.1:

- (i) $\overset{1}{\text{C}}\text{H}_2 = \overset{2}{\text{C}} = \text{O}$
C-1 is sp^2 hybridised.
C-2 is sp hybridised.
- (ii) $\overset{1}{\text{C}}\text{H}_3 - \overset{2}{\text{C}}\text{H} = \overset{3}{\text{C}}\text{H}_2$
C-1 is sp^3 hybridised.
C-2 is sp^2 hybridised.
C-3 is sp^2 hybridised.
- (iii) $\overset{1}{\text{C}}\text{H}_3 - \overset{2}{\text{C}} = \overset{\text{O}}{\parallel} - \overset{3}{\text{C}}\text{H}_3$
C-1 is sp^3 hybridised.
- (iv) $\overset{1}{\text{C}}\text{H}_2 = \overset{2}{\text{C}}\text{H} - \overset{3}{\text{C}} = \text{N}$
C-2 is sp^2 hybridised.
C-3 is sp hybridised.
- (v) C_6H_6
All the 6 carbon atoms in benzene are sp^2 hybridised.

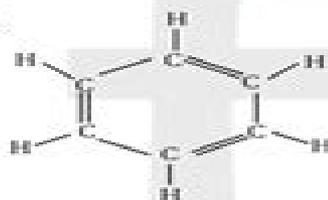
Question 8.2:

Indicate the σ and π bonds in the following molecules :

C_6H_6 , C_6H_{12} , CH_2Cl_2 , $\text{CH}_2 = \text{C} = \text{CH}_2$, CH_3NO_2 , HCONHCH_3

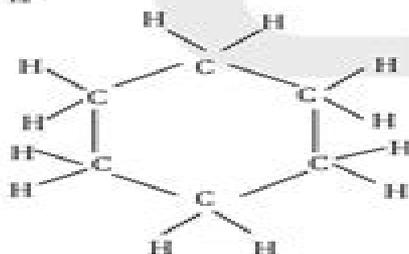
Answer 8.2:

(i) C_6H_6



There are six C-C sigma ($\sigma_{\text{C-C}}$) bonds, six C-H sigma ($\sigma_{\text{C-H}}$) bonds and three C=C pi ($\pi_{\text{C=C}}$) resonating bonds in this compound.

(ii) C_6H_{12}



There are six C-C sigma ($\sigma_{\text{C-C}}$) bonds and twelve C-H sigma ($\sigma_{\text{C-H}}$) bonds in this compound.

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Table of Contents Chemistry Chapter 8

1. Understanding the eBook Chemistry Chapter 8
 - The Rise of Digital Reading Chemistry Chapter 8
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 8
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Chemistry Chapter 8
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 8
 - Personalized Recommendations
 - Chemistry Chapter 8 User Reviews and Ratings
 - Chemistry Chapter 8 and Bestseller Lists
5. Accessing Chemistry Chapter 8 Free and Paid eBooks
 - Chemistry Chapter 8 Public Domain eBooks
 - Chemistry Chapter 8 eBook Subscription Services
 - Chemistry Chapter 8 Budget-Friendly Options
6. Navigating Chemistry Chapter 8 eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 8 Compatibility with Devices
 - Chemistry Chapter 8 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 8
 - Highlighting and Note-Taking Chemistry Chapter 8
 - Interactive Elements Chemistry Chapter 8
8. Staying Engaged with Chemistry Chapter 8

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 8
9. Balancing eBooks and Physical Books Chemistry Chapter 8
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 8
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 8
- Setting Reading Goals Chemistry Chapter 8
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 8
- Fact-Checking eBook Content of Chemistry Chapter 8
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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