

Chemistry 338 2nd Semester Final Exam Review Sheet

Unit 1 - Matter & Measurement – Chapters 1 & 2

1. A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

$$\% \text{ error} = \frac{\text{experimental} - \text{accepted}}{\text{accepted}} \times 100$$

$$\% \text{ error} = \frac{4.2 - 5.0}{5.0} \times 100 = -16\% \text{ error}$$

Unit 2 – Periodic Table – Chapter 5

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Wungsten ion	¹⁸⁴ ₇₄ W ⁺²	74	74	110	72	184	+2

3. Using the following atomic radius values, predict the atomic radius of potassium (K):
 Li – 145 pm, Na – 180 pm, K - ?, Rb – 235 pm
 K – 220 pm

Unit 3 – Reactions/Compounds – Chapters 7 & 8

4. Classify the compound as ionic or covalent. Name the compound
- Fe₂O₃ - iron (III) oxide (ionic - roman numerals for transition metals, change ending to -ide)
 - Cl₂O₈ - dichlorine octaoxide (covalent - use prefixes, change ending to -ide)
5. Classify the compound as ionic or covalent. Write the formula
- Trisulfur Dinitride – S₃N₂ (covalent)
 - Zinc Hydroxide – Zn(OH)₂ (ionic - look up ion charges and criss-cross) Zn⁺² OH⁻¹
 - Chromium (III) nitrate – Cr(NO₃)₃ (ionic - criss-cross) Cr⁺³ NO₃⁻²
6. Write the equation, balance it, and classify the type of reaction:
 Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate
- $$1 \text{ FeSO}_4 + 2 \text{ NaOH} \rightarrow 1 \text{ Fe(OH)}_2 + 1 \text{ Na}_2\text{SO}_4$$
- double displacement (two ionic compounds reacting)
7. Balance the equation and classify the type of reaction:
- $$1 \text{ Cu} + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Cu(NO}_3)_2 + 2 \text{ Ag}$$
- Single displacement (element and a compound reacting)

Unit 4 – Moles – Chapter 3

8. Calculate the molar mass of Fe₂(SO₄)₃.
- $$\begin{aligned} \text{Fe (2)(55.9 g/mol)} &= 111.8 \text{ g/mol} \\ \text{S (3)(32.1 g/mol)} &= 96.3 \text{ g/mol} \\ + \text{O (12)(16.0 g/mol)} &= +192 \text{ g/mol} \\ \hline &= 400.1 \text{ g/mol} \end{aligned}$$

Chemistry Spring Semester Final Exam Review Sheet

Diane M. Bunce, Cinzia M. Muzzi



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Survival Handbook for the New Chemistry Instructor Diane M. Bunce, Cinzia M. Muzzi, 2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching Serving as a reference to answer specific questions new chemistry faculty encounter this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching

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