

f-block elements  $\div$  Group - 8, Period  $\begin{cases} 6 - \text{lanthanoids} \\ 7 - \text{actinoids} \end{cases}$

$\rightarrow$  are those in

which the 4f and 5f orbitals are progressively filled

$\rightarrow$  Placed below the table in two long periods (6<sup>th</sup>, 7<sup>th</sup>)

$\rightarrow$  Formal members of group 8, but taken out

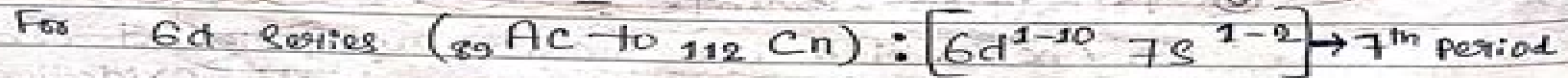
$\rightarrow$  Inner transition metals

Lanthanoids and Actinoids

## # The Transition Elements $\div$ (d-block)

The very name transition given to the elements of d-block is only because of their position between s and p block elements

$\rightarrow$  General E.C is



# Inorganic Chemistry Notes For Iit Jee

**SJ Ball**



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