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THE RIGHT INSULATION MATTERS

# INSULATION HANDBOOK

## Part 2: Professional Installation Guide - Version 5

Insulation installation for ceilings, walls & floors

An independent publication of the  
Insulation Council of Australia and New Zealand

# Statoil Insulation Handbook

**RC Schank**



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was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key 4 ko<sub>2</sub> s 2 h 2 o l 4 koh aq 3 o<sub>2</sub> g if a reaction vessel contains 0 mol ko<sub>2</sub> and 0 mol h 2 o what is the limiting reactant

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole h<sub>2</sub>s and multiply it with 2 mole h<sub>2</sub>o over 2 mole h<sub>2</sub>s from the equation  

$$8 3 \text{ mole h}_2\text{s} \times \frac{2 \text{ mole h}_2\text{o}}{2 \text{ mole h}_2\text{s}} = 8 3 \text{ mole h}_2\text{o}$$
  
 8 3 mole h<sub>2</sub>s x 18g h<sub>2</sub>o i mole h<sub>2</sub>o 149 4g h<sub>2</sub>o is your theoretical yield so percent yield 137 1g h<sub>2</sub>o actual yield 149 4g h<sub>2</sub>o x 100  
 91 77 problems 6 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
  
 if the reaction of 21 3 grams of o<sub>2</sub> produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have o 2 16 2 32 grams mole 21 3 grams 0 66 moles 32 g m

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calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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 $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  26.982 g  $\text{Al} \times 4 \text{ mol Al} \times 1 \text{ mol Al}_2\text{O}_3$  c determine the percent yield  
yield actual  $\times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$  theoretical 302.3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be  
weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12  
c is the limiting reagent 4.33 g of h<sub>2</sub> are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s  
remains when the reaction is complete

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent  
yield theoretical  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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 $\text{CH}_3\text{OH}$  can be produced by the following reaction  $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of co is  
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