

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 12

## STUDY GUIDE FOR CONTENT MASTERY

## Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data for table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of  $\text{O}_2$ ,  $\text{H}_2$ , and  $\text{H}_2\text{O}$  are 32.00 g/mol, 2.016 g/mol, and 18.02 g/mol, respectively.

Amount of $\text{O}_2$	Amount of $\text{H}_2$	Amount of $\text{H}_2\text{O}$	Limiting Reactant	Amount of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	$\text{H}_2$	2 molecules $\text{O}_2$
2 molecules	6 molecules	4 molecules	$\text{O}_2$	2 molecules $\text{H}_2$
1.00 mol	2.00 mol	2.00 mol	none	none
2.00 mol	4.00 mol	4.00 mol	$\text{H}_2$	2.00 mol $\text{O}_2$
3.00 mol	6.00 mol	6.00 mol	$\text{H}_2$	3.00 mol $\text{O}_2$
4.00 mol	8.00 mol	8.00 mol	$\text{H}_2$	4.00 mol $\text{O}_2$
5.00 mol	10.00 mol	10.00 mol	$\text{H}_2$	5.00 mol $\text{O}_2$
6.00 mol	12.00 mol	12.00 mol	$\text{H}_2$	6.00 mol $\text{O}_2$
7.00 mol	14.00 mol	14.00 mol	$\text{H}_2$	7.00 mol $\text{O}_2$
8.00 mol	16.00 mol	16.00 mol	$\text{H}_2$	8.00 mol $\text{O}_2$
9.00 mol	18.00 mol	18.00 mol	$\text{H}_2$	9.00 mol $\text{O}_2$
10.00 mol	20.00 mol	20.00 mol	$\text{H}_2$	10.00 mol $\text{O}_2$
11.00 mol	22.00 mol	22.00 mol	$\text{H}_2$	11.00 mol $\text{O}_2$
12.00 mol	24.00 mol	24.00 mol	$\text{H}_2$	12.00 mol $\text{O}_2$
13.00 mol	26.00 mol	26.00 mol	$\text{H}_2$	13.00 mol $\text{O}_2$
14.00 mol	28.00 mol	28.00 mol	$\text{H}_2$	14.00 mol $\text{O}_2$
15.00 mol	30.00 mol	30.00 mol	$\text{H}_2$	15.00 mol $\text{O}_2$
16.00 mol	32.00 mol	32.00 mol	$\text{H}_2$	16.00 mol $\text{O}_2$
17.00 mol	34.00 mol	34.00 mol	$\text{H}_2$	17.00 mol $\text{O}_2$
18.00 mol	36.00 mol	36.00 mol	$\text{H}_2$	18.00 mol $\text{O}_2$
19.00 mol	38.00 mol	38.00 mol	$\text{H}_2$	19.00 mol $\text{O}_2$
20.00 mol	40.00 mol	40.00 mol	$\text{H}_2$	20.00 mol $\text{O}_2$

balanced equation mole ratio = 2 mol  $\text{H}_2\text{O}$ /1 mol  $\text{O}_2$

16.00 g  $\text{O}_2$  = 1 mol  $\text{O}_2$ /32.00 g  $\text{O}_2$  = 0.500 mol  $\text{O}_2$

20.00 g  $\text{H}_2$  = 1 mol  $\text{H}_2$ /2.016 g  $\text{H}_2$  = 9.920 mol  $\text{H}_2$

actual mole ratio = 9.920 mol  $\text{H}_2$ /0.500 mol  $\text{O}_2$  = 19.84 mol  $\text{H}_2$ /1 mol  $\text{O}_2$

Because the actual mole ratio of  $\text{H}_2$  to  $\text{O}_2$  is larger than the balanced equation mole ratio of  $\text{H}_2$  to  $\text{O}_2$ , there is an excess of  $\text{H}_2$ ;  $\text{O}_2$  is the limiting reactant.

Mass of  $\text{H}_2$  used = 0.500 mol  $\text{O}_2$   $\times$  2 mol  $\text{H}_2$ /1 mol  $\text{O}_2$  = 1.000 mol  $\text{H}_2$

0.500 mol  $\text{O}_2$   $\times$  32.00 g  $\text{O}_2$ /1 mol  $\text{O}_2$  = 16.00 g  $\text{O}_2$

Mass of  $\text{H}_2\text{O}$  produced = 0.500 mol  $\text{O}_2$   $\times$  2 mol  $\text{H}_2\text{O}$ /1 mol  $\text{O}_2$  = 1.000 mol  $\text{H}_2\text{O}$

Excess  $\text{H}_2$  = 20.00 g  $\text{H}_2$  - 1.000 g  $\text{H}_2$  = 19.00 g  $\text{H}_2$

Excess  $\text{O}_2$  = 20.00 g  $\text{O}_2$  - 16.00 g  $\text{O}_2$  = 4.00 g  $\text{O}_2$

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## CHAPTER 12

## STUDY GUIDE FOR CONTENT MASTERY

## Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



**Example Problem:** The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and  $\text{Ga}_3\text{N}_2$  are 69.72 g/mol and 177.48 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield volume = 7.42 g  $\text{Ga}_3\text{N}_2$
2. mass of reactant = 7.00 g Ga
3. number of moles of reactant = 7.00 g Ga  $\times$   $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$  = 7.00 g Ga = 1 mol Ga/69.72 g Ga
4. number of moles of product = 0.05 mol Ga  $\times$   $\frac{2 \text{ mol Ga}_3\text{N}_2}{1 \text{ mol Ga}}$  = 0.10 mol  $\text{Ga}_3\text{N}_2$
5. theoretical yield = 0.0500 mol  $\text{Ga}_3\text{N}_2$   $\times$   $\frac{177.48 \text{ g Ga}_3\text{N}_2}{1 \text{ mol Ga}_3\text{N}_2}$  = correct
6. percent yield =  $\frac{7.42 \text{ g Ga}_3\text{N}_2}{177.48 \text{ g Ga}_3\text{N}_2} \times 100$  =  $\frac{7.42 \text{ g Ga}_3\text{N}_2}{177.48 \text{ g Ga}_3\text{N}_2} \times 100$

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