

# Chemistry

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(Chapter 1)(Solutions) (Exercises)  
(Class 12)

## Question 1.1:

Define the term solution. How many types of solutions are formed? Write briefly about each type with an example.

## Answer 1.1:

A homogeneous mixture of two or more non-reacting substances whose composition varies within certain fixed limits, known as a solution. Generally, the component that is present in the largest quantity is known as *solvent*. Solvent determines the physical state in which solution exists. One or more components present in the solution other than solvent are called solutes.

When the solutions contains two, three or four components, it is known as binary, ternary or quaternary solutions respectively. On the basis of physical state of components, the solutions are of the following types:

- 1) When both solute and solvent are in solid state. Example: Alloys.
- 2) When solute is in solid and solvent is in liquid state. Example: Sugar or salt solutions.
- 3) When solute is in solid and solvent is in gaseous state. Example: iodine vapours in air.
- 4) When both solute and solvent are in liquid state. Example: alcohol in water.
- 5) When solute is in liquid and solvent is in solid state. Example: Zinc amalgam.
- 6) When solute is in liquid and solvent is in gaseous state. Example: water vapour in air.
- 7) When both solute and solvent are in gaseous state. Example: air.
- 8) When solute is in gaseous and solvent is in liquid state. Example: aerated drinks.
- 9) When solute is in gaseous and solvent is in solid state. Example: dissolve gas in mineral.

## Question 1.2:

Give an example of a solid solution in which the solute is a gas.

## Answer 1.2:

Dissolved gases in minerals.

## Question 1.3:

Define the following terms:

- (i) Mole fraction                      (ii) Molality                      (iii) Molarity                      (iv) Mass percentage.

## Answer 1.3:

(i) Mole fraction: Commonly used symbol for mole fraction is  $x$  and subscript used on the right hand side of  $x$  denotes the component. It is defined as:

$$\text{Mole fraction of a component} = \frac{\text{Number of moles of the component}}{\text{Number of moles of all the component}}$$

(ii) Molality: Molality ( $m$ ) is defined as the number of moles of the solute per kilogram (kg) of the solvent and is expressed as:

$$\text{Molality (m)} = \frac{\text{Moles of solute}}{\text{Mass of solvent in kg}}$$

(iii) Molarity: Molarity ( $M$ ) is defined as number of moles of solute dissolved in one litre (or one cubic decimetre) of solution.

$$\text{Molarity (M)} = \frac{\text{Moles of solute}}{\text{Volume of solution in litre}}$$

(iv) Mass percentage: The mass percentage of a component of a solution is defined as:

$$\text{Mass \% of a component} = \frac{\text{Mass of the component in the solution}}{\text{Total mass of the solution}} \times 100$$

# Chemistry Chapter 1

**EW Minium**



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*Scribner's Monthly* Josiah Gilbert Holland,Richard Watson Gilder,1914

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### **Table of Contents Chemistry Chapter 1**

1. Understanding the eBook Chemistry Chapter 1
  - The Rise of Digital Reading Chemistry Chapter 1
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 1
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 1
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 1
  - Personalized Recommendations
  - Chemistry Chapter 1 User Reviews and Ratings

- Chemistry Chapter 1 and Bestseller Lists
- 5. Accessing Chemistry Chapter 1 Free and Paid eBooks
  - Chemistry Chapter 1 Public Domain eBooks
  - Chemistry Chapter 1 eBook Subscription Services
  - Chemistry Chapter 1 Budget-Friendly Options
- 6. Navigating Chemistry Chapter 1 eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 1 Compatibility with Devices
  - Chemistry Chapter 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 1
  - Highlighting and Note-Taking Chemistry Chapter 1
  - Interactive Elements Chemistry Chapter 1
- 8. Staying Engaged with Chemistry Chapter 1
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 1
- 9. Balancing eBooks and Physical Books Chemistry Chapter 1
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 1
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 1
  - Setting Reading Goals Chemistry Chapter 1
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 1
  - Fact-Checking eBook Content of Chemistry Chapter 1
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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