

AP Chemistry: Chapter 14 Practice Test

10. Given the following acids and K_a values:

HClO_4	HOAc	HCN	HF
-----------------	---------------	--------------	-------------

1×10^7	1.76×10^{-5}	4.93×10^{-10}	3.53×10^{-4}
-----------------	-----------------------	------------------------	-----------------------

which shows the conjugate bases listed by increasing strength?

- | | |
|--|--|
| [A] ClO_4^- , F^- , OAc^- , CN^- | [B] CN^- , F^- , OAc^- , ClO_4^- |
| [C] CN^- , OAc^- , F^- , ClO_4^- | [D] CN^- , ClO_4^- , F^- , OAc^- |
| [E] ClO_4^- , OAc^- , CN^- , F^- | |
15. Given that the K_a for HOCl is 3.5×10^{-8} , calculate the K value for the reaction of HOCl with OH^- .
- | | | |
|---------------------------|-----------------------|--------------------------|
| [A] 3.5×10^{-22} | [B] 3.5×10^6 | [C] 2.9×10^{-7} |
| [D] 3.5×10^{-8} | [E] none of these | |
32. Acetic acid ($\text{HC}_2\text{H}_3\text{O}_2$) is a weak acid ($K_a = 1.8 \times 10^{-5}$). Calculate the pH of a 17.6 M $\text{HC}_2\text{H}_3\text{O}_2$ solution.
- | | | | | |
|---------|---------|---------|----------|---------|
| [A] 4.3 | [B] 6.4 | [C] 1.7 | [D] 0.97 | [E] 7.4 |
|---------|---------|---------|----------|---------|
52. What concentration of acetic acid ($K_a = 1.80 \times 10^{-5}$) has the same pH as that of 5.00×10^{-3} M HCl ?
- | | | |
|-----------------------------|-----------------------------|------------|
| [A] 5.00×10^{-3} M | [B] 3.60×10^{-3} M | [C] 1.39 M |
| [D] 5.00 M | [E] none of these | |

Consider the reaction $\text{HOCl} + \text{F}^- \rightleftharpoons \text{HF} + \text{OCl}^-$

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The Princeton Review



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