

16. ^{137}I has a half-life of 8.04 days. Assuming you start with a 1.25 mg sample of ^{137}I , how much will remain after 13.8 days?

- A) 0.625 mg
 B) 0.268 mg
 C) 0.422 mg
 D) 0.380 mg

$$k = \frac{\ln 2}{t_{1/2}} = \frac{\ln 2}{8.04 \text{ days}} = 0.0862 \text{ days}^{-1}$$

$$\ln(N_t) = -kt + \ln(N_0)$$

$$\ln(N_t) = -(0.0862 \text{ days}^{-1})(13.8 \text{ days}) + \ln(1.25 \text{ mg}) = -0.823$$

$$N_t = e^{-0.823} = \mathbf{0.440 \text{ mg}}$$

17. Which one of the following devices converts radioactive emissions to light for detection?

- A) Geiger counter
 B) photomultiplier tube
 C) scintillation counter
 D) none of these

18. Which one of the following is true?

- A) Most spontaneous nuclear reactions are exothermic.
 B) Some spontaneous nuclear reactions are endothermic.
 C) All spontaneous nuclear reactions are exothermic.
 D) There is no relationship between exothermicity and spontaneity in nuclear reactions.

19. Binding energy per nucleon is greatest for nuclei with...

- A) large mass numbers
 B) intermediate mass numbers
 C) small mass numbers

20. When two atoms of ^{235}U are fused to form one atom of ^{238}U , the total energy involved is $3.83 \times 10^{-10} \text{ J}$. What is the total change in mass for this reaction? ($c = 3.00 \times 10^8 \text{ m/s}$)

- A) $4.26 \times 10^{-17} \text{ g}$
 B) $4.26 \times 10^{-16} \text{ g}$
 C) $4.26 \times 10^{-18} \text{ g}$
 D) $4.26 \times 10^{-19} \text{ g}$

$$E = mc^2 \quad \Rightarrow \quad m = \frac{E}{c^2} = \frac{3.83 \times 10^{-10} \text{ J}}{(3.00 \times 10^8 \text{ m/s})^2} = 4.26 \times 10^{-16} \text{ kg} = \left(\frac{1000 \text{ g}}{1 \text{ kg}}\right) = \mathbf{4.26 \times 10^{-16} \text{ g}}$$

Study Nuclear Chemistry Answer Key

RC Schank



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This book delves into Study Nuclear Chemistry Answer Key. Study Nuclear Chemistry Answer Key is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Study Nuclear Chemistry Answer Key, encompassing both the fundamentals and more intricate discussions.

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