

CHAPTER 4 REVIEW

Arrangement of Electrons in Atoms**SECTION 2****SHORT ANSWER** Answer the following questions in the space provided.

- How many quantum numbers are used to describe the properties of electrons in atomic orbitals?
(a) 1 (b) 2 (c) 3 (d) 4
- A spherical electron cloud surrounding an atomic nucleus would best represent
(a) an *s*-orbital. (b) a *p*-orbital.
(c) a combination of two different *p*-orbitals. (d) a combination of an *s* and a *p*-orbital.
- How many electrons can an energy level of $n = 4$ hold?
(a) 32 (b) 24 (c) 8 (d) 6
- How many electrons can an energy level of $n = 2$ hold?
(a) 32 (b) 24 (c) 8 (d) 6
- Compared with an electron for which $n = 2$, an electron for which $n = 4$ has more
(a) spin. (b) particle nature.
(c) energy. (d) wave nature.
- According to Bohr, which is the point in the figure below where electrons cannot reside?
(a) point A (b) point B (c) point C (d) point D



- According to the quantum theory, point D in the above figure represents
(a) the fixed position of an electron.
(b) the furthest position from the nucleus that an electron can achieve.
(c) a position where an electron probably exists.
(d) a position where an electron cannot exist.

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B Lingard



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