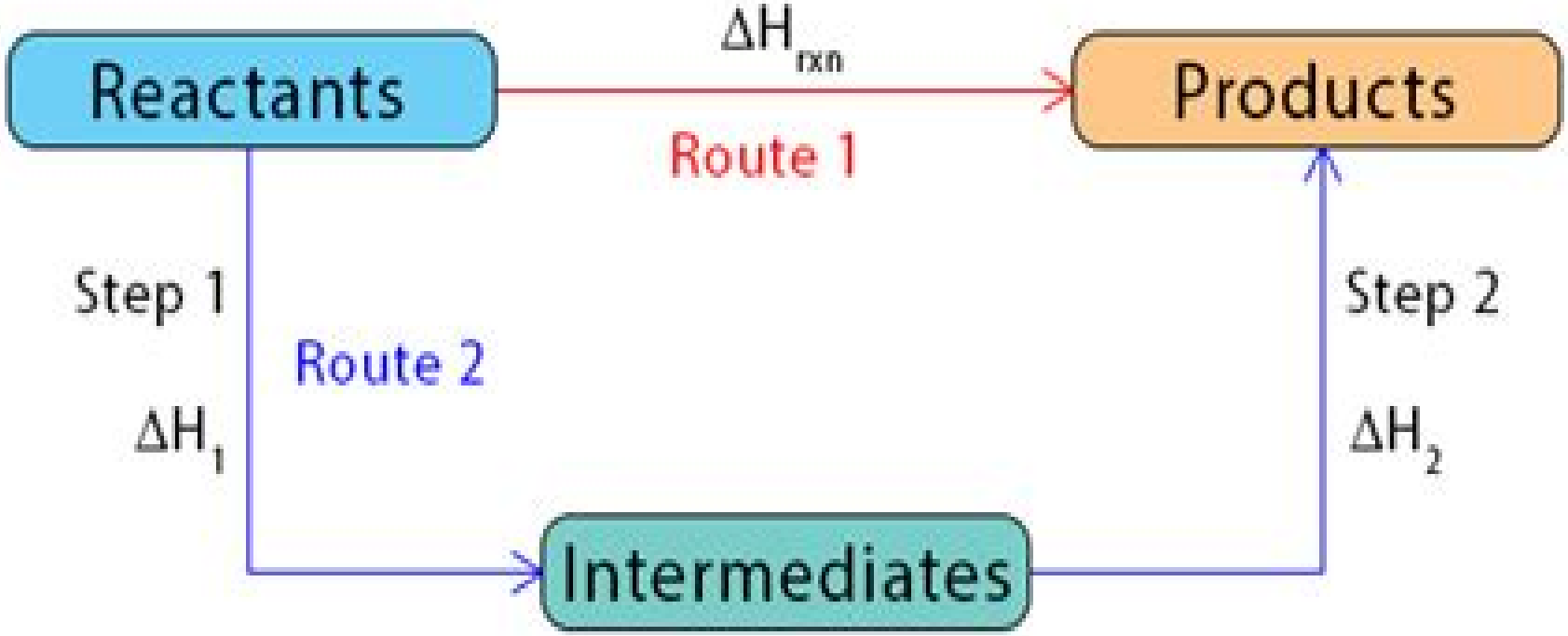
Hess's Law

Total enthalpy change (ΔH) of a reaction equals to the sum of all the enthalpy changes occurring in each step of the reaction



$$\Delta H_{rxn} = \Delta H_1 + \Delta H_2$$

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