

## Chapter 2 - Practice Test #1

7. An equilibrium system shifts left when the
- A. rate of the forward reaction is equal to the rate of the reverse reaction.
  - B. rate of the forward reaction is less than the rate of the reverse reaction.
  - C. rate of the forward reaction is greater than the rate of the reverse reaction.
  - D. rate of the forward reaction and the reverse reaction are constant.
8. A 1.00 L flask contains a gaseous equilibrium system. The addition of reactants to this flask results in a
- A. shift left and a decrease in the concentration of the products.
  - B. shift left and a increase in the concentration of the products.
  - C. shift right and a decrease in the concentration of the products.
  - D. shift right and a increase in the concentration of the products.
9. Consider the following equilibrium:  $\text{CH}_4\text{(g)} + \text{H}_2\text{O(g)} + \text{heat} = \text{CO(g)} + 3\text{H}_2\text{(g)}$   
In which of the following will both stresses shift the equilibrium to the right?
- A. a decrease in temperature and a decrease in volume
  - B. a increase in temperature and a decrease in volume
  - C. a decrease in temperature and a increase in volume
  - D. a increase in temperature and a increase in volume
10. Consider the following equilibrium:  $2\text{SO}_2\text{(g)} + \text{O}_2\text{(g)} = 2\text{SO}_3\text{(g)}$   $\Delta H = -198 \text{ kJ}$   
There will be no shift in this equilibrium when
- A. more  $\text{O}_2$  is added.
  - B. a catalyst is added.
  - C. the volume is increased.
  - D. the temperature is increased.
11. Consider the following equilibrium:  $2\text{Fe(s)} + 3\text{H}_2\text{O(g)} = \text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{(g)}$   
The equilibrium expression is
- A.  $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][\text{H}_2]^3}{[\text{Fe}]^2[\text{H}_2\text{O}]^3}$
  - B.  $\text{Keq} = \frac{[\text{Fe}_2\text{O}_3][3\text{H}_2]}{[2\text{Fe}][3\text{H}_2\text{O}]}$
  - C.  $\text{Keq} = \frac{[\text{H}_2]^3}{[\text{H}_2\text{O}]^3}$
  - D.  $\text{Keq} = [\text{H}_2]^3$
12. Consider the following equilibrium:  $\text{N}_2\text{O}_4\text{(g)} = 2\text{NO}_2\text{(g)}$   $\text{Keq} = 0.133$   
At equilibrium, the  $[\text{N}_2\text{O}_4]$  is equal to

# Chemical Equilibrium Practice Test Answers

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