

Matter and Change, pp. 1-4

- | | | |
|----|----|---|
| 2. | 5. | c |
| 3. | 5. | c |
| 4. | 7. | c |
| 9. | 3. | a |
10. mass and volume
 11. group or family
 12. homogeneous
 13. products
 14. metal
 15. solid state
 16. chemical properties
 17. atom
 18. compound
 19. nonmetal
 20. mixture
 21. pure substance
 22. basic
 23. metalloid
 24. reactants: carbon and oxygen; product: carbon dioxide
 25. reactant: mercury(II) oxide; products: mercury and oxygen
 26. f
 27. e
 28. c
 29. a
 30. b
 31. i
 32. d
 33. h
 34. e
 35. physical change
 36. chemical change
 37. physical change
 38. physical change
 39. chemical change
 40. physical change
 41. In a solid, particles are packed together in relatively fixed positions. The particles vibrate about a fixed point. In a liquid, particles are close together but can flow around one another. In a gas, particles are at great distances from one another, compared to the particles of liquids and solids.
 42. The composition of a pure substance is the same throughout and does not vary from sample to sample. Pure substances include elements and compounds. A mixture contains more than one substance and can vary in composition from one sample to another and sometimes within different parts of the same sample.
 43. Although energy can be absorbed or released in a

change, it is not destroyed or created. It simply takes a different form.

43. Homogeneous mixtures are uniform in composition; heterogeneous mixtures are not.
44. An element cannot be broken down, but water can be decomposed into hydrogen and oxygen by passing an electrical current through it (electrolysis).

Measurements and Calculations, pp. 5-8

- | | |
|------|------|
| a | 5. c |
| 2. c | 6. c |
| 3. c | 7. a |
| 4. a | 8. a |
9. time
 10. mass
 11. density
 12. energy
 13. length or distance
 14. volume
 15. area
 16. qualitative
 17. quantitative
 18. qualitative
 19. quantitative
 20. 3.00×10^7 km/s
 21. three
 22. 0.026 g or 2.6×10^{-2} g
 23. 2.5×10^{-4} L
 24. quantity
 25. derived units
 26. conversion factor
 27. 0.0432 kg or 4.32×10^{-2} kg
 28. 5400 mL or 5.4×10^3 mL
 29. 300. K
 30. 1.05×10^3 J
 31. 3.51×10^3 cm
 32. A hypothesis is a testable statement that can be used to make predictions and to carry out further experiments. A theory is a broad generalization that explains a body of facts or phenomena.
 33. Mass is a measure of the quantity of matter. Weight is a measure of the gravitational pull on matter. Mass does not depend on gravitational attraction.
 34. Models are visual, verbal, or mathematical representations. They are used to explain how phenomena occur or how data or events are related.
 35. The graph of two quantities that are directly proportional is a straight line through the origin.

Chemistry Matter Change Chapter 16 Answer Key

William L. Ramsey



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Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change *Ebook: Chemistry: The Molecular Nature of Matter and Change* Silberberg, 2015-01-16
Ebook Chemistry The Molecular Nature of Matter and Change Chemistry McGraw-Hill Staff, 2001-03 *Chemistry* Martin Stuart Silberberg, 1996
This general chemistry text centres on the theme that observable change in chemical systems is the result of molecular change The aims of this edition are to enable students to perceive matter and change at the molecular level and to help build student confidence in their ability to solve chemical problems as they discover the relevance of chemistry to their lives

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Everything in the universe regardless of its size shape color or physical state is made up of matter **General Chemistry** Ralph H. Petrucci, 1989 Introduction to Physics and Chemistry Bill W. Tillery, 1992

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