

KEY

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RESULTS

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1

Chemistry Lab Experiment

Determination of the Molar Volume of a Gas

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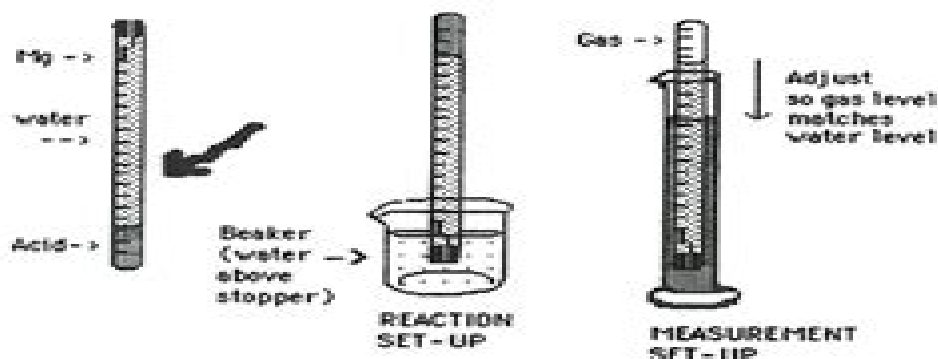
Name:

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Purpose:

- To experimentally determine the molar volume of hydrogen gas and to compare it with the $\frac{22.4\text{-L}}{\text{mole}}$ theoretical STP value.
- To relate gas law concepts learned in MA Frameworks Topic 6 (e.g., the Combined Gas Law and Dalton's Law of Partial Pressures) to previously learned concepts involving balancing chemical equations and stoichiometric relationships.

Introduction:



The **water displacement method** can be used to capture the hydrogen gas liberated when a strip of magnesium metal reacts with a 3.0-*M* solution of hydrochloric acid. The total gas volume of H₂O vapor and H₂ adjusted to the room's barometric pressure can be determined from the **eudiometer** (graduated gas-measuring tube) by partially immersing the tube into a large container of water until the liquid level inside the tube matches the water level in the large container.

According to **Dalton's Law of Partial Pressures**, the total gas pressure in the eudiometer is the sum of the pressure of the hydrogen gas collected plus the **water vapor pressure**. Symbolically, $P_{\text{total}} = P_{\text{H}_2} + P_{\text{H}_2\text{O}}$. Subtracting the water vapor pressure from the total pressure gives the pressure of the dry hydrogen gas.

Chemistry Molar Volume Of Hydrogen Lab Answers

Terry C. Jones



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Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

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