



Carbon tetrachloride ( $\text{CCl}_4$ )

$\mu = 0\text{D}$

$\text{CCl}_4$  is a symmetrical molecule. Therefore, the dipole moments of all four C–Cl bonds cancel each other. Hence, its resultant dipole moment is zero.

As shown in the above figure, in  $\text{CHCl}_3$ , the resultant of dipole moments of two C–Cl bonds is opposed by the resultant of dipole moments of one C–H bond and one C–Cl bond. Since the resultant of one C–H bond and one C–Cl bond dipole moments is smaller than two C–Cl bonds, the opposition is to a small extent. As a result,  $\text{CHCl}_3$  has a small dipole moment of 1.08 D.

On the other hand, in case of  $\text{CH}_2\text{Cl}_2$ , the resultant of the dipole moments of two C–Cl bonds is strengthened by the resultant of the dipole moments of two C–H bonds. As a result,  $\text{CH}_2\text{Cl}_2$  has a higher dipole moment of 1.60 D than  $\text{CHCl}_3$  i.e.,  $\text{CH}_2\text{Cl}_2$  has the highest dipole moment.

Hence, the given compounds can be arranged in the increasing order of their dipole moments as:



Question 10.5:

A hydrocarbon  $\text{C}_5\text{H}_{10}$  does not react with chlorine in dark but gives a single monochloro compound  $\text{C}_5\text{H}_9\text{Cl}$  in bright sunlight. Identify the hydrocarbon.

Answer

A hydrocarbon with the molecular formula,  $\text{C}_5\text{H}_{10}$ , belongs to the group with a general molecular formula  $\text{C}_n\text{H}_{2n}$ . Therefore, it may either be an alkene or a cycloalkane.

Since hydrocarbon does not react with chlorine in the dark, it cannot be an alkene. Thus, it should be a cycloalkane.

Further, the hydrocarbon gives a single monochloro compound,  $\text{C}_5\text{H}_9\text{Cl}$  by reacting with chlorine in bright sunlight. Since a single monochloro compound is formed, the hydrocarbon must contain H-atoms that are all equivalent. Also, as all H-atoms of a cycloalkane are equivalent, the hydrocarbon must be a cycloalkane. Hence, the said compound is cyclopentane.



Cyclopentane ( $\text{C}_5\text{H}_{10}$ )

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Test Banner



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# Chapter 3 Chemistry Test Answers

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### **Table of Contents Chapter 3 Chemistry Test Answers**

1. Understanding the eBook Chapter 3 Chemistry Test Answers
  - The Rise of Digital Reading Chapter 3 Chemistry Test Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 3 Chemistry Test Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 3 Chemistry Test Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 3 Chemistry Test Answers
  - Personalized Recommendations
  - Chapter 3 Chemistry Test Answers User Reviews and Ratings
  - Chapter 3 Chemistry Test Answers and Bestseller Lists
5. Accessing Chapter 3 Chemistry Test Answers Free and Paid eBooks
  - Chapter 3 Chemistry Test Answers Public Domain eBooks
  - Chapter 3 Chemistry Test Answers eBook Subscription Services
  - Chapter 3 Chemistry Test Answers Budget-Friendly Options

6. Navigating Chapter 3 Chemistry Test Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 3 Chemistry Test Answers Compatibility with Devices
  - Chapter 3 Chemistry Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 3 Chemistry Test Answers
  - Highlighting and Note-Taking Chapter 3 Chemistry Test Answers
  - Interactive Elements Chapter 3 Chemistry Test Answers
8. Staying Engaged with Chapter 3 Chemistry Test Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 3 Chemistry Test Answers
9. Balancing eBooks and Physical Books Chapter 3 Chemistry Test Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 3 Chemistry Test Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 3 Chemistry Test Answers
  - Setting Reading Goals Chapter 3 Chemistry Test Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 3 Chemistry Test Answers
  - Fact-Checking eBook Content of Chapter 3 Chemistry Test Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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