

Chapter 10 mole concept advance additional notes

What is mole?









This is an atom of carbon-12.
It is chosen as the standard atom.
Its A_r is taken as 12. Then other atoms are compared with it.



This is a magnesium atom.
It is twice as heavy as a carbon-12 atom, so its A_r is 24.
So it follows that ...

A mole of a substance is defined as the number of atoms in an element

More examples of moles

<p>Sodium is made of single sodium atoms. Its symbol is Na. Its A_r is 23.</p> 	<p>Iodine is made of iodine molecules. Its formula is I_2. Its M_r is 254.</p> 	<p>Water is made of water molecules. Its formula is H_2O. Its M_r is 18.</p> 
 <p>This is 23 grams of sodium. It contains 6.02×10^{23} sodium atoms, or 1 mole of sodium atoms.</p>	 <p>This is 254 grams of iodine. It contains 6.02×10^{23} iodine molecules, or 1 mole of iodine molecules.</p>	 <p>The beaker contains 18 grams of water, or 6.02×10^{23} water molecules, or 1 mole of water molecules.</p>

Finding the mass of a mole

In order to find the mass of any substance, you must follow these steps:

- Write down the symbol or formula of the substance
- Find its M_r or A_r .
- Express that mass in grams(g)

Examples:

Substance	Symbol or formula	A_r	M_r	Mass of 1 mole
helium	He	He = 4	mass as single atoms	4 grams
oxygen	O_2	O = 16	$2 \times 16 = 32$	32 grams
ethanol	C_2H_5OH	C = 12 H = 1 O = 16	$2 \times 12 = 24$ $6 \times 1 = 6$ $1 \times 16 = 16$ <hr/> 46	46 grams

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Table of Contents Chem11 The Mole Concept Notes 10 Mrowen

1. Understanding the eBook Chem11 The Mole Concept Notes 10 Mrowen
 - The Rise of Digital Reading Chem11 The Mole Concept Notes 10 Mrowen
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem11 The Mole Concept Notes 10 Mrowen
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem11 The Mole Concept Notes 10 Mrowen
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem11 The Mole Concept Notes 10 Mrowen
 - Personalized Recommendations
 - Chem11 The Mole Concept Notes 10 Mrowen User Reviews and Ratings

- Chem11 The Mole Concept Notes 10 Mrowen and Bestseller Lists
- 5. Accessing Chem11 The Mole Concept Notes 10 Mrowen Free and Paid eBooks
 - Chem11 The Mole Concept Notes 10 Mrowen Public Domain eBooks
 - Chem11 The Mole Concept Notes 10 Mrowen eBook Subscription Services
 - Chem11 The Mole Concept Notes 10 Mrowen Budget-Friendly Options
- 6. Navigating Chem11 The Mole Concept Notes 10 Mrowen eBook Formats
 - ePub, PDF, MOBI, and More
 - Chem11 The Mole Concept Notes 10 Mrowen Compatibility with Devices
 - Chem11 The Mole Concept Notes 10 Mrowen Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem11 The Mole Concept Notes 10 Mrowen
 - Highlighting and Note-Taking Chem11 The Mole Concept Notes 10 Mrowen
 - Interactive Elements Chem11 The Mole Concept Notes 10 Mrowen
- 8. Staying Engaged with Chem11 The Mole Concept Notes 10 Mrowen
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem11 The Mole Concept Notes 10 Mrowen
- 9. Balancing eBooks and Physical Books Chem11 The Mole Concept Notes 10 Mrowen
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem11 The Mole Concept Notes 10 Mrowen
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chem11 The Mole Concept Notes 10 Mrowen
 - Setting Reading Goals Chem11 The Mole Concept Notes 10 Mrowen
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chem11 The Mole Concept Notes 10 Mrowen
 - Fact-Checking eBook Content of Chem11 The Mole Concept Notes 10 Mrowen
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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