

QUALITATIVE ANALYSIS

Experiment	Observation	Inference
<u>Preliminary Test</u>		
1. Noted the colour of the given salt	1. Blue or Bluish Green 2. Greenish 3. Light Green 4. Dark Brown 5. Light Pink 6. Colourless	Cu^{2+} or Ni^{2+} Ni^{2+} Fe^{2+} Fe^{3+} Mn^{2+} Absence of Cu^{+2} , Ni^{2+} , Fe^{2+} , Fe^{3+} , Mn^{2+}
2. Noted the solubility	1. Soluble in water 2. Soluble in dilute HCl	NH_4^+ may be present Absence of group I cations
3. Dry Heating Test: Heated a pinch of the salt in a dry Test tube	1. Colour, odourless gas 2. Colourless gas with rotten egg smell 3. Colourless gas with vinegar smell 4. Colourless gas with ammonia smell 5. Reddish brown gas 6. Greenish yellow gas 7. Violet gas	CO_3^{2-} may be present S^{2-} may be present CH_3COO^- may be present NH_4^+ may be present NO_3^- or Br^- may be present Cl^- may be present I^- may be present
<u>Identification of acid radical (Anion)</u>		
1. Dilute Acid Test: Treated a pinch of the substance With dilute H_2SO_4	1. Colourless, odourless gas with brisk effervescence which turned lime water milky. 2. Vinegar Smell 3. No characteristic Reaction	CO_3^{2-} present $\text{CO}_3^{2-} + 2\text{H}^+ \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$ CH_3COO^- Present $\text{CH}_3\text{COO}^- + \text{H}^+ \rightarrow \text{CH}_3\text{COOH}$ Absence of CO_3^{2-} , CH_3COO^-
2. Concentrated Acid Test: Treated a pinch of the substance With concentrated H_2SO_4	1. Colourless gas with pungent Smell which gave dense white fumes when a glass rod dipped in NH_4OH was brought near the mouth of the test tube 2. Reddish brown vapours with Pungent smell. Solution turns red	Cl^- present $\text{Cl}^- + \text{H}^+ \rightarrow \text{HCl}$ $\text{HCl} + \text{NH}_4\text{OH} \rightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$ Br^- present $\text{Br}^- + \text{H}^+ \rightarrow \text{HBr}$ $2\text{HBr} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Br}_2 + \text{SO}_2$
	3. Deep violet vapours with pungent smell. Sides of test tube	I^- present $\text{I}^- + \text{H}^+ \rightarrow \text{HI}$ $2\text{HI} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{I}_2 + \text{SO}_2$

Chemistry Practical Qualitative Analysis Sheet

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A Laboratory Manual of Organic Chemistry for Beginners Arnold Frederick Holleman, 1913 Manual of Qualitative Chemical Analysis C. Remigius Fresenius, 1890 **The Annual Catalogue of Purdue University, Lafayette, Indiana ... with Announcements for ...** Purdue University, 1920 Outlines of Theoretical Chemistry Frederick Hutton Getman, 1913

The Educational Times, and Journal of the College of Preceptors ,1898 **Elementary Linear Perspective of Form and Shadow** Samuel Edward Warren, 1904 The Medical Times and Gazette ,1873 **Annual Announcement of the School of Pharmacy** Purdue University. School of Pharmacy, 1926 **Pharmaceutical Journal** ,1900

Engineering ,1906 **The pharmaceutical journal and transactions** ,1872 **Elementary instruction in chemical analysis [tr. from Anleitung zur qualitativen chemischen Analyse]** ed. by J.L. Bullock Carl Remigius Fresenius, 1872 **Dental Brief** Thomas Bromwell Welch, Wilbur F. Litch, Alfred P. Lee, 1907 *A systematic handbook of volumetric analysis or, The quantitative estimation of chemical substances by measure* Francis Sutton, 1871 **Merck's Report** Theodore Weicker, 1902 *American Horticultural Manual: Comprising the leading principles and practices connected with the propagation, culture, and improvement of fruits, nuts, ornamental trees, shrubs, and plants in the United States and Canada* Joseph L. Budd, Niels Ebbesen Hansen, 1902 **Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain), 1895 Calendar for the Year ,1901 *Laboratory Manual of Elementary Chemistry* Jabez Montgomery, Roy Burnett Smith, 1904 **The Chemical News and Journal of Physical Science** ,1904

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