

CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

Labels for the periodic table:
 H — ATOMIC NUMBER
 1.008 — ATOMIC WEIGHT
 H — ATOMIC SYMBOL



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For neutral atoms, $Z = +$ the number of electrons. The mass number, A , is the sum of protons and neutrons in the nucleus. Isotopes of an element have the same Z , different A . The atomic weight is the weighted average of isotopes. The atomic weight is the weighted average of isotopes for a given element.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number, n , is the number of shells. $n = 1, 2, 3, \dots$
 Azimuthal quantum number, l , is the number of subshells. $l = 0, 1, 2, \dots, n-1$
 Magnetic quantum number, m_l , is the number of orbitals in a subshell. $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$
 Spin quantum number, m_s , is the number of electrons in an orbital. $m_s = +\frac{1}{2}, -\frac{1}{2}$

NUCLEAR CHEMISTRY

Nuclear reactions occur when nucleons (protons and neutrons) are added or removed. Nuclear energy is released or absorbed. The mass defect is the difference between the mass of the nucleus and the sum of the masses of its constituent nucleons. The binding energy is the energy required to separate the nucleus into its constituent nucleons.

MANY-ELECTRON ATOMS

Electrons and protons in an atom are held together by the electrostatic force. The Pauli Exclusion Principle states that no two electrons in an atom can have the same set of quantum numbers. The Aufbau Principle states that electrons fill orbitals in order of increasing energy.

n	l	m _l	m _s	Orbitals
1	0	0	$+\frac{1}{2}, -\frac{1}{2}$	1s
2	0	0	$+\frac{1}{2}, -\frac{1}{2}$	2s
2	1	0, ± 1	$+\frac{1}{2}, -\frac{1}{2}$	2p
3	0	0	$+\frac{1}{2}, -\frac{1}{2}$	3s
3	1	0, ± 1	$+\frac{1}{2}, -\frac{1}{2}$	3p
3	2	0, $\pm 1, \pm 2$	$+\frac{1}{2}, -\frac{1}{2}$	3d

Atomic configuration: The arrangement of electrons in an atom. The Pauli Exclusion Principle, the Aufbau Principle, and the Hund's Rule of Maximum Multiplicity are used to determine the atomic configuration. The ground state configuration is the most stable configuration.



Block	Group	Element	Symbol	Atomic Number		
s-block	1	H	H	1		
		Li	Li	3		
	2	Be	Be	4		
		Mg	Mg	12		
	p-block	13	B	B	5	
			Al	Al	13	
		14	C	C	6	
			Si	Si	14	
		15	N	N	7	
			P	P	15	
		16	O	O	8	
			S	S	16	
		17	F	F	9	
			Cl	Cl	17	
		18	Ne	Ne	10	
			Ar	Ar	18	
		d-block	3	Sc	Sc	21
				Ti	Ti	22
4	V		V	23		
	Cr		Cr	24		
5	Mn		Mn	25		
	Fe		Fe	26		
6	Co		Co	27		
	Ni		Ni	28		
7	Cu		Cu	29		
	Zn		Zn	30		
f-block	8	La	La	57		
		Ce	Ce	58		
	9	Pr	Pr	59		
		Nd	Nd	60		
	10	Pm	Pm	61		
		Sm	Sm	62		
	11	Eu	Eu	63		
		Gd	Gd	64		
	12	Tb	Tb	65		
		Dy	Dy	66		
13	Ho	Ho	67			
	Er	Er	68			
14	Tm	Tm	69			
	Yb	Yb	70			
15	Lu	Lu	71			
	Hf	Hf	72			

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Marco Cascella



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