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Simulation: Ionic and Covalent Bonding

Background

In this investigation you will bond select atoms. Based upon the types of atoms that you choose to combine, you will create either an ionic compound or a covalent compound. You will have the opportunity to analyze the differences between these different types of compounds and to predict the number of atoms needed to create each, as well as learn how to appropriately name them.

1. Describe the difference between an atom and a molecule:

Atoms are single particles, and molecules are made up of multiple atoms

2. Where are metal atoms located on the periodic table? Where are non-metal atoms located on the periodic table?

Metals are on the left side and non-metals are on the right

3. What subatomic particle(s) participate in chemical bonding?

electron

4. In your own words, define valence electron.

electrons on outside ring which allow for bonding, negative charge

5. How can you determine the number of valence electrons in an atom using the periodic table?

Look at the atomic #

6. Draw a Lewis Dot Structure for the following atoms:

a. Strontium (Sr)



b. Carbon (C)



c. Iodine (I)



d. Xenon (Xe)

*Check your answers before moving on to the next portion of the activity.

Ionic Bonding Lab Answers

JA Banks



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Table of Contents Ionic Bonding Lab Answers

1. Understanding the eBook Ionic Bonding Lab Answers
 - The Rise of Digital Reading Ionic Bonding Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonding Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonding Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonding Lab Answers
 - Personalized Recommendations

- Ionic Bonding Lab Answers User Reviews and Ratings
- Ionic Bonding Lab Answers and Bestseller Lists
- 5. Accessing Ionic Bonding Lab Answers Free and Paid eBooks
 - Ionic Bonding Lab Answers Public Domain eBooks
 - Ionic Bonding Lab Answers eBook Subscription Services
 - Ionic Bonding Lab Answers Budget-Friendly Options
- 6. Navigating Ionic Bonding Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Bonding Lab Answers Compatibility with Devices
 - Ionic Bonding Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonding Lab Answers
 - Highlighting and Note-Taking Ionic Bonding Lab Answers
 - Interactive Elements Ionic Bonding Lab Answers
- 8. Staying Engaged with Ionic Bonding Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonding Lab Answers
- 9. Balancing eBooks and Physical Books Ionic Bonding Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonding Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Bonding Lab Answers
 - Setting Reading Goals Ionic Bonding Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Bonding Lab Answers
 - Fact-Checking eBook Content of Ionic Bonding Lab Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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