

1-6, 8, 9, 11

## Hydrolysis of Salts and pH of Buffer Solutions

## Pre-lab Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. Define Brønsted-Lowry acids and bases.

According to Brønsted-Lowry definitions, an acid is a proton donor and a base is a proton acceptor.

2

- Which of the following ions will react with water in a hydrolysis reaction:

$K^+$ ,  $Ba^{2+}$ ,  $Cu^{2+}$ ,  $Zn^{2+}$ ,  $F^-$ ,  $SO_3^{2-}$ , and  $Cl^-$ ?

$Cu^{2+}$ ,  $Zn^{2+}$ ,  $F^-$  and  $SO_3^{2-}$  will undergo hydrolysis.

3

- For those ions in question 2 that undergo hydrolysis, write net ionic equations for the hydrolysis reaction.

Example:  $Cu^{2+}(aq) + 2H_2O(l) \rightleftharpoons Cu(OH)_2(s) + 2H^+(aq)$ .

4

- The  $K_a$  for HCN is  $4.9 \times 10^{-10}$ . What is the value of  $K_b$  for  $CN^-$ ?

$$K_b = \frac{1.0 \times 10^{-14}}{4.9 \times 10^{-10}} = 2.0 \times 10^{-5}$$

5. What are the conjugate base and conjugate acid of  $H_2PO_4^-$ ?

The conjugate base is  $HPO_4^{2-}$  and the conjugate acid is  $H_3PO_4$ .

6

- From what acid and what base were the following salts made:  $K_2SO_4$ ,  $NH_4Cl$ , and  $Ca(NO_3)_2$ ?

For  $Ca^{2+}$  and  $SO_4^{2-}$  salt is made from  $Ca(OH)_2$  and  $H_2SO_4$ .

7

- Define the term *salt*.

# Ph Properties Of Buffer Solutions Answer Key Pre Lab

**CH Cherryholmes**



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