

## Exercise 2: Creating an Activity Series

Data Table 2. Redox Reactions of Copper, Lead, and Zinc.

Redox Metal	Redox ID	Solution	Immediate Observations	30-Minute Observations
Cu	0	$Pb(NO_3)_2$	No noticeable changes	None
	0	$Zn(NO_3)_2$	No noticeable changes	None
Pb	0	$CuSO_4$	Blue solution, a rust color appeared on the surface	Blue color faded and appeared to be rusty
	0	$ZnSO_4$	None	None
Zn	0	$CuSO_4$	Appeared to be dark, rusty, and fuzzy appearance, blue solution	Appeared to be rusted, was a dark colored almost black, less blue
	0	$Pb(NO_3)_2$	Dark almost black, clear solution	Appeared to be fuzzy, dark, and grey colored, clear solution

Data Table 3. Potential Redox Reactions and Chemical Equations.

Metal and Aqueous Solution	Reaction Occurred?	Chemical Equation
$Cu + Pb(NO_3)_2$	No	No reaction
$Cu + Zn(NO_3)_2$	No	No reaction
$Pb + CuSO_4$	Yes	$Pb + CuSO_4 \rightarrow Cu + PbSO_4$
$Pb + Zn(NO_3)_2$	No	$Pb + Zn(NO_3)_2 \rightarrow Zn + Pb(NO_3)_2$
$Zn + CuSO_4$	Yes	$Zn + CuSO_4 \rightarrow Cu + ZnSO_4$
$Zn + Pb(NO_3)_2$	Yes	$Zn + Pb(NO_3)_2 \rightarrow Pb + Zn(NO_3)_2$

# Activity Series Lab Report Answers

**VM Jensen**



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