

STOICHIOMETRY PROBLEMS

1. The combustion of a sample of butane, C_4H_{10} (lighter fluid), produced 2.46 grams of water.



- (a) How many moles of water formed?
- (b) How many moles of butane burned?
- (c) How many grams of butane burned?
- (d) How much oxygen was used up in moles?
- (e) How much oxygen was used up in grams?

2. Sodium metal burns in air according to the balanced reaction below.



- (a) How many moles of oxygen are needed to react with 9.5 g of sodium completely?
- (b) How many grams of sodium are needed to produce 12.5 g of sodium oxide?

3. In the combustion of 54.5 g of butane (C_4H_{10}), how many grams of CO_2 are produced? Write and balance the equation before solving.

Guided Practice Problems 11 Stoichiometry

Kathleen Armour



Guided Practice Problems 11 Stoichiometry:

Decoding **Guided Practice Problems 11 Stoichiometry**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Guided Practice Problems 11 Stoichiometry**," a mesmerizing literary creation penned by a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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