

Chapter 2: The Chemistry of Life

State Standards

Central Concept: Chemical elements form organic molecules that interact to perform the basic functions of life.

- 1.1 Recognize that biological organisms are composed primarily of many few elements. The six most common are Carbon, Hydrogen, Nitrogen, Oxygen, Phosphorus, and Sulfur.
- 1.2 Describe the basic molecular structures and primary functions of the four major categories of organic molecules (carbohydrates, lipids, proteins, and nucleic acids).
- 1.3 Explain the role of enzymes as catalysts that lower the activation energy of biochemical reactions. Identify factors, such as pH and temperature, that have an effect on enzymes.

Key Terms

Atom	Adhesion	Monosaccharide
Nucleus	Mixture	Lipid
Proton, Neutron, Electron	Solution	Nucleic Acid
Element	Solute	Nucleotide
Isotope	Solvent	Protein
Compound	Suspension	Amino Acid
Ionic Bond	pH Scale	Chemical Reaction
Ion	Acid	Reactant
Covalent Bond	Base	Product
Molecule	Buffer	Activation Energy
Van der Waals Forces	Monomer	Catalyst
Hydrogen Bond	Polymer	Enzyme
Cohesion	Carbohydrate	Substrate

You should be able to:

1. Identify the three subatomic particles found in atoms.
2. Explain how all of the isotopes of an element are similar and how they are different.
3. Explain how compounds are different from their component elements.
4. Describe the two main types of chemical bonds.
5. Discuss the unique properties of water.
6. Differentiate between solutions and suspensions.
7. Explain what acidic solutions and basic solutions are.
8. Describe the unique qualities of carbon.
9. Describe the structures and functions of each of the four groups of macromolecules.
10. Explain how chemical reactions affect chemical bonds.
11. Describe how energy changes affect how easily a chemical reaction will occur.
12. Explain why enzymes are important to living things.

Biology Chapter 2 The Chemistry Of Life Study Guide

William K. Purves



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