

## Chemical Bonding

### SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one  
(a) formula unit. (c) crystal.  
(b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of  
(a) molecules. (c) dipoles.  
(b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is  
(a) higher in potential energy. (c) equal in potential energy.  
(b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?  
(a) Compound A is not an ionic compound.  
(b) It will be more difficult to break the bonds in compound A than those in compound B.  
(c) Compound B has larger crystals than compound A.  
(d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally  
(a) stronger than the attractive forces among formula units in ionic bonding.  
(b) weaker than the attractive forces among formula units in ionic bonding.  
(c) approximately equal to the attractive forces among formula units in ionic bonding.  
(d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

**Very strong overall attraction between positive and negative charges**

- What type of energy best represents the strength of an ionic bond?

**Lattice energy**

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

# Modern Chemistry Chapter 6 Worksheet Answers

**RC Schank**



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## **Table of Contents Modern Chemistry Chapter 6 Worksheet Answers**

1. Understanding the eBook Modern Chemistry Chapter 6 Worksheet Answers
  - The Rise of Digital Reading Modern Chemistry Chapter 6 Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 6 Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Chapter 6 Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 6 Worksheet Answers
  - Personalized Recommendations
  - Modern Chemistry Chapter 6 Worksheet Answers User Reviews and Ratings
  - Modern Chemistry Chapter 6 Worksheet Answers and Bestseller Lists

5. Accessing Modern Chemistry Chapter 6 Worksheet Answers Free and Paid eBooks
  - Modern Chemistry Chapter 6 Worksheet Answers Public Domain eBooks
  - Modern Chemistry Chapter 6 Worksheet Answers eBook Subscription Services
  - Modern Chemistry Chapter 6 Worksheet Answers Budget-Friendly Options
6. Navigating Modern Chemistry Chapter 6 Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Chemistry Chapter 6 Worksheet Answers Compatibility with Devices
  - Modern Chemistry Chapter 6 Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 6 Worksheet Answers
  - Highlighting and Note-Taking Modern Chemistry Chapter 6 Worksheet Answers
  - Interactive Elements Modern Chemistry Chapter 6 Worksheet Answers
8. Staying Engaged with Modern Chemistry Chapter 6 Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Chapter 6 Worksheet Answers
9. Balancing eBooks and Physical Books Modern Chemistry Chapter 6 Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Chapter 6 Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter 6 Worksheet Answers
  - Setting Reading Goals Modern Chemistry Chapter 6 Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Chapter 6 Worksheet Answers
  - Fact-Checking eBook Content of Modern Chemistry Chapter 6 Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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