

Vol. of 1 mol of ideal gas at STP/NTP = 22.4 L

उसका सही अंश
- 3 same sabse kam significant figure hoga pt. ke bad utne hi hoga ans me.

*** Significant fig**
 0 to the left of the 1st non 0 digit are sig. not sig.
 eg. 0.0012 \Rightarrow 2
 Terminal 0 are not sig. if there is no decimal pt.
 eg. 1000 \Rightarrow 1 fig.
 0 at end is sig. if there is decimal pt.
 Exponential term are not considered sig.

अंश - 3/1/31
 $a \times b = c$ } - add/subst. is least amount
 \Rightarrow 3 same sabse kam fig hoga vo sig. fig ans me hoga

*** LAW OF MASS CONSV**
 In a reactⁿ \rightarrow actual mass of reactant = mass of prod.
 Rasm me zidna use hua.

Exceptⁿ \rightarrow Nuclear Rxn (E = mc²)
 PppWedⁿ \rightarrow Balanced Rxn

MOLE CONCEPT

*** LAW OF CONST PROPORTⁿ**
 Elements in a comp are in fixed ratio by mass.
 eg. CO₂ \rightarrow 3:8
 12 32g
 % of C = $\frac{3}{11} \times 100$

Q. cal. the mass of C req. to form 132g of CO₂.
 \Rightarrow C + 2O \rightarrow CO₂
 12g 32g 44g
 44g CO₂ ke lie 12g C
 \therefore 132g ke lie = $\frac{12}{44} \times 132 = 36\text{g}$

*** Law of Multiple Propⁿ**
 elements should be same or same fixed mass
 eg. CO, CO₂ 1:2 ratio
 Other eg. N₂O, N₂O₃, N₂O₅
 exceptⁿ \rightarrow N₂O₃, N₂O₅

*** Avogadro's Law**
 Same (P, T, R, V) of molecules of gases same (mass eg. 1 mol of)
 X₂, X₃, X₄
 2 : 3 : 4 molecules
 2 : 3 : 4 atoms [different]

*** Gay Lussac's Law**
 CONDINS: +T, P same
 at least 2 gases.
 H₂ + Cl₂ \rightarrow 2HCl
 1 : 1 : 2
 Vol. R₂ U₂

*** DALTON'S ATOMIC THEORY**
 Atoms \rightarrow Neither created nor destroyed

*** MOLE CONCEPT**
 $\frac{A}{Z} \times (P+N)$
 Atomic No. (P)
 Mass of Proton = $1.67 \times 10^{-27} \text{ kg}$
 Neutron = $1.67 \times 10^{-27} \text{ kg}$
 Electron = $9.1 \times 10^{-31} \text{ kg}$
 MP $\approx 1837 \text{ MeV}$

*** Atomic mass unit (amu)**
 $1 \text{ amu} = \frac{1}{12} \times \text{Mass of } ^{12}\text{C}$
 $1 \text{ amu} = 1.67 \times 10^{-27} \text{ kg}$
 eg. Mass of $^{12}\text{C} = 12 \text{ amu}$
*** Relative Atomic mass**
 Unitless
 eg. mass. A.M. of C = 12

*** Molecular Mass**
 H₂SO₄ = 98 amu
 \rightarrow Mass of 1 molecule of H₂SO₄

*** FORMULA MASS**: for ionic comp.
*** MOLE**
 Amount of substance.
 1 mole = No particles (6.022 x 10²³)
 1 amu $\leq 1.67 \times 10^{-27} \text{ kg}$
 Na = 6.022 x 10²³
 1 amu y Na $\leq 1\text{g}$

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