

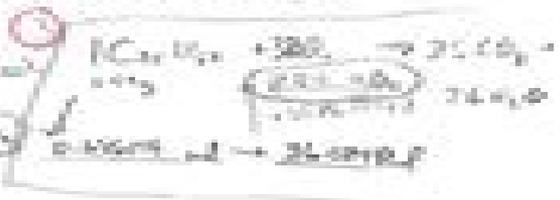
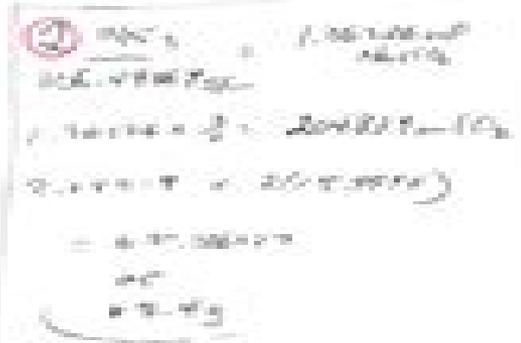
KEY

1. Complete and classify the following partial equations

- a) $1Fe + 3H_2SO_4 \rightarrow 1FeSO_4 + 3H_2$ (single repl)
- b) $2Na + Cl_2 \rightarrow 2NaCl$
- c) $1Zn + 1Pb(NO_3)_2 \rightarrow 1Zn(NO_3)_2 + 1Pb$ (single repl)
- d) $Al(OH)_3 + HCl \rightarrow AlCl_3 + 3H_2O$ (acid base)
- e) $2Cu + O_2 \rightarrow 2CuO$ (synthesis)
- f) $4Cu + 1O_2 \rightarrow 2Cu_2O$ (synthesis)
- g) $Al_2O_3 \rightarrow 2Al + 3/2O_2$ (decomposition)
- h) $1N_2 + 3H_2 \rightarrow 2NH_3$ (synthesis)
- i) $C_2H_6 + O_2 \rightarrow 2CO_2 + 3H_2O$ (combustion)
- j) $1Mg + 2HNO_3 \rightarrow 1Mg(NO_3)_2 + 1H_2$ (single repl)
- k) $Na + H_2O \rightarrow NaOH + H_2$ (single repl)
- l) $3NaOH + H_3PO_4 \rightarrow 1Na_3PO_4 + 3H_2O$ (acid base)
- m) $H_2SO_4 + Al(OH)_3 \rightarrow Al_2(SO_4)_3 + 6H_2O$ (acid base)
- n) $2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O$ (combustion)
- o) $AlCl_3 \rightarrow Al + 3/2Cl_2$ (decomposition)
- p) $2KOH + H_2SO_4 \rightarrow 1K_2SO_4 + 2H_2O$ (acid base)
- q) $C_2H_6 + O_2 \rightarrow 2CO_2 + 3H_2O$ (combustion)
- r) $1NaOH + HCl \rightarrow 1NaCl + 1H_2O$ (acid base)
- s) $NaHCO_3 + H_2SO_4 \rightarrow Na_2SO_4 + 2H_2O + 2CO_2$ (acid base)
- t) $1CH_4 + 2O_2 \rightarrow 1CO_2 + 2H_2O$ (combustion)
- u) $H_2O \rightarrow H_2 + 1/2O_2$ (decomposition)
- v) $2HNO_3 + CuCO_3 \rightarrow 1Cu(NO_3)_2 + 1H_2O + 1CO_2$ (acid base)
- w) $CH_4 + O_2 \rightarrow CO_2 + 2H_2O$ (combustion)
- x) $Pb + H_2SO_4 \rightarrow 1PbSO_4 + 1H_2$ (single repl)
- y) $KI + Br_2 \rightarrow 2KBr + I_2$ (single repl)
- z) $1Ba(OH)_2 + 2HNO_3 \rightarrow 1Ba(NO_3)_2 + 2H_2O$ (acid base)
- aa) $Pb + Cu(NO_3)_2 \rightarrow Pb(NO_3)_2 + Cu$ (single repl)



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2. Given: $2NaClO_3 \rightarrow 2NaCl + 3O_2$ calculate the mass of oxygen gas produced if 143 g of $NaClO_3$ is decomposed

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($C_{21}H_{44}$)

254g

4. If 1.40 moles of butane (C_4H_{10}) is burned

- a) calculate the number of moles of oxygen gas required (9.8 mol O_2)
- b) calculate the number of moles of carbon dioxide produced (5.6 mol CO_2)
- c) calculate the mass of water produced (10.8 mol H_2O)

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate

- a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. (27.4g Al)
- b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. (192 g Zn)

6. Given: $Cu + 2AgNO_3 \rightarrow Cu(NO_3)_2 + 2Ag$

- a) How many grams of Cu are required to react with 3.00 g of silver nitrate? (0.544g Cu)
- b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? (1.70g Ag)

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Michael Brown



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