

# 8

## Acids, bases and salts

### Key objectives

By the end of this section, you should be able to:

- describe the characteristic reactions of acids
  - state that bases are oxides or hydroxides of metals
  - state that alkalis are soluble bases
  - describe the characteristic reactions of bases
  - state that aqueous solutions of acids contain  $H^+$  ions and aqueous solutions of alkalis contain  $OH^-$  ions
  - describe how to use pH, as measured with universal indicator paper, to compare hydrogen ion concentration, neutrality, relative acidity and alkalinity
  - describe the neutralisation reaction
- define acids and bases in terms of proton transfer
  - explain the difference between weak and strong acids in terms of dissociation
- classify oxides as acidic, basic or amphoteric, with examples
  - describe the general solubility rules for salts
  - describe the preparation of soluble salts by the reaction of an acid with:
    - an alkali
    - excess metal, insoluble base or carbonate
- describe the preparation of insoluble salts by precipitation
- define a hydrated substance and an anhydrous substance
- define the term water of crystallisation in crystals, including  $CuSO_4 \cdot 5H_2O$  and  $CoCl_2 \cdot 6H_2O$

### Key terms

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Term	Definition
<b>Acid</b>	A substance which dissolves in water, producing $H^+(aq)$ ions as the only positive ion. <b>A proton (<math>H^+</math>) donor.</b>
<b>Alkali</b>	A soluble base which produces $OH^-(aq)$ ions in water.
<b>Anhydrous salt</b>	A salt which has had its water of crystallisation removed.
<b>Base</b>	A substance which neutralises an acid, producing a salt and water as the only products. Bases are oxides or hydroxides of metals. (Ammonia is also a base.) <b>A proton (<math>H^+</math>) acceptor.</b>
<b>Indicator</b>	A substance that shows whether a substance is acidic or alkaline by changing colour.
<b>Neutralisation</b>	The process in which an acid reacts with a base to form water.
<b>pH scale</b>	A scale running from 0 to 14 used to express the acidity or alkalinity of a substance.
<b>Saturated solution</b>	A solution containing the maximum amount of dissolved solute in the solvent at a given temperature.
<b>Water of crystallisation</b>	Water incorporated into the structure of a substance as it crystallises, for example in copper(II) sulfate pentahydrate ( $CuSO_4 \cdot 5H_2O$ ).

### 8.1 Acids and alkalis

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#### Acids

**Acids** are substances that produce  $H^+$  ions when they are dissolved in water.

Acids are defined as proton ( $H^+$ ) donors.

# Chapter 8 Acids Bases And Ph Quia

**M Lipman**



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## **Table of Contents Chapter 8 Acids Bases And Ph Quia**

1. Understanding the eBook Chapter 8 Acids Bases And Ph Quia
  - The Rise of Digital Reading Chapter 8 Acids Bases And Ph Quia
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Acids Bases And Ph Quia
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 8 Acids Bases And Ph Quia
  - User-Friendly Interface

4. Exploring eBook Recommendations from Chapter 8 Acids Bases And Ph Quia
  - Personalized Recommendations
  - Chapter 8 Acids Bases And Ph Quia User Reviews and Ratings
  - Chapter 8 Acids Bases And Ph Quia and Bestseller Lists
5. Accessing Chapter 8 Acids Bases And Ph Quia Free and Paid eBooks
  - Chapter 8 Acids Bases And Ph Quia Public Domain eBooks
  - Chapter 8 Acids Bases And Ph Quia eBook Subscription Services
  - Chapter 8 Acids Bases And Ph Quia Budget-Friendly Options
6. Navigating Chapter 8 Acids Bases And Ph Quia eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 8 Acids Bases And Ph Quia Compatibility with Devices
  - Chapter 8 Acids Bases And Ph Quia Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 8 Acids Bases And Ph Quia
  - Highlighting and Note-Taking Chapter 8 Acids Bases And Ph Quia
  - Interactive Elements Chapter 8 Acids Bases And Ph Quia
8. Staying Engaged with Chapter 8 Acids Bases And Ph Quia
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 8 Acids Bases And Ph Quia
9. Balancing eBooks and Physical Books Chapter 8 Acids Bases And Ph Quia
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 8 Acids Bases And Ph Quia
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 8 Acids Bases And Ph Quia
  - Setting Reading Goals Chapter 8 Acids Bases And Ph Quia
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chapter 8 Acids Bases And Ph Quia
  - Fact-Checking eBook Content of Chapter 8 Acids Bases And Ph Quia
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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