

CHAPTER 4 REVIEW

Chemical Equations and Reactions

Short Review

Write a balanced chemical equation for the reaction described.

1. _____
 - a. balanced chemical equation represents all the following except:
 - i an experimentally established law
 - ii substances by which elements combine to form products
 - iii identities of reactant and products in a chemical reaction
 - iv relative quantities of reactant and products in a chemical reaction
 - b. According to the law of conservation of mass, the total mass of the
 - i reactants is _____
 - i a) always more than the total mass of the products
 - ii) always less than total mass of the products
 - iii) sometimes more and sometimes less than the total mass of the products
 - iv) always equal to total mass of the products
2. Write a balanced chemical equation for the following chemical reactions and label. For each reaction that will occur, identify the reaction type and describe the chemical equation using its four products. Comment regarding the final product. (Dimensionality rules can be found in Figure 4.3 of the text)
 - a. $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow$

 - b. $\text{H}_2\text{O} + \text{NaCl} \rightarrow$

 - c. $\text{H}_2\text{O} + \text{NaOH} \rightarrow$

 - d. $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow$

Chapter 8 Chemical Reactions Guided Reading Answers

M Woodhall



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