



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - H_2SO_4
 - H_2O_2
- Write the formula for the following molecular compounds:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

- Write the molecular formula for each of the following compounds:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide
- Write the molecular formula for each of the following compounds:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride: $\text{H}-\text{F}$
- Draw the Lewis structure for methane: C with four H atoms bonded to it.
- Draw the Lewis structure for ethane: $\text{H}-\text{C}-\text{C}-\text{H}$ with appropriate H atoms.
- Draw the Lewis structure for acetylene: $\text{H}-\text{C}\equiv\text{C}-\text{H}$

- Draw the Lewis structure for each of the following:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

SECTION 8.3 BONDING: THE ORBITALS

- Write the hybridization for each of the following:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide

Chapter 8 Covalent Bonding Test B Answers

Terry C. Jones



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