

## CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron (II) sulfide, the roman numeral tells us (2)
  - How many atoms of Fe are in one formula unit.
  - How many sulfide ions can be attached to the iron atom.
  - The charge on each Fe ion.
  - The total positive charge of the formula unit.
- The result of changing a subscript in a correctly written chemical formula is (1)
  - Change the number of ions represented in the formula.
  - Change the charges on the other ions in the compound.
  - Change the formula so that it no longer represents the compound.
  - Have no effect on the formula.
- Derive a generalization for when an acid name will end in the suffix -ic or -ous. If ending in -ic, the anion ended in -ate. If ending in -ous, the anion ended in -ite.
- Derive a generalization for when an acid name will begin with the prefix hydro-. If prefix is hydro-, the anion ended in -ite. This is almost always a monatomic ion.
- Fill in the blanks in the table below.

Compound Name	Formula
Acetic acid	$\text{CH}_3\text{CO}_2\text{H}$
Aluminum sulfate	$\text{Al}_2(\text{SO}_4)_3$
Lead II carbonate	$\text{PbCO}_3$
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Sulfuric acid	$\text{H}_2\text{SO}_4$
Nitrogen pentoxide	$\text{N}_2\text{O}_5$
Copper II carbonate	$\text{CuCO}_3$
Sodium sulfate	$\text{Na}_2\text{SO}_4$
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Mercury sulfate	$\text{HgSO}_4$
Hydroic acid	$\text{HI}$
Magnesium perchlorate	$\text{Mg}(\text{ClO}_4)_2$
Iron II nitrate	$\text{Fe}(\text{NO}_3)_2$
Iron III nitrate	$\text{Fe}(\text{NO}_3)_3$
Carbon Dioxide	$\text{CO}_2$
Sulfur dioxide	$\text{SO}_2$
Sulfur pentachloride	$\text{SCl}_5$
Tetraphosphorus trisulfide	$\text{P}_4\text{S}_3$
Carbonic acid	$\text{H}_2\text{CO}_3$

- Name the following compound and identify its molecular ions or acid.
  - $\text{Ca}(\text{NO}_3)_2$ , calcium (II) nitrate - II
  - $\text{Fe}(\text{NO}_3)_3$ , Iron (III) nitrate - III
  - $\text{Fe}(\text{OH})_2$ , Iron (II) hydroxide - II
  - $\text{N}_2\text{O}_5$ , Nitrogen pentoxide - III

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## **Table of Contents Chapter 9 Chemical Names Formulas Answers Page 221**

1. Understanding the eBook Chapter 9 Chemical Names Formulas Answers Page 221
  - The Rise of Digital Reading Chapter 9 Chemical Names Formulas Answers Page 221
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Chemical Names Formulas Answers Page 221
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 9 Chemical Names Formulas Answers Page 221
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Chemical Names Formulas Answers Page 221
  - Personalized Recommendations

- Chapter 9 Chemical Names Formulas Answers Page 221 User Reviews and Ratings
- Chapter 9 Chemical Names Formulas Answers Page 221 and Bestseller Lists
- 5. Accessing Chapter 9 Chemical Names Formulas Answers Page 221 Free and Paid eBooks
  - Chapter 9 Chemical Names Formulas Answers Page 221 Public Domain eBooks
  - Chapter 9 Chemical Names Formulas Answers Page 221 eBook Subscription Services
  - Chapter 9 Chemical Names Formulas Answers Page 221 Budget-Friendly Options
- 6. Navigating Chapter 9 Chemical Names Formulas Answers Page 221 eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 9 Chemical Names Formulas Answers Page 221 Compatibility with Devices
  - Chapter 9 Chemical Names Formulas Answers Page 221 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 9 Chemical Names Formulas Answers Page 221
  - Highlighting and Note-Taking Chapter 9 Chemical Names Formulas Answers Page 221
  - Interactive Elements Chapter 9 Chemical Names Formulas Answers Page 221
- 8. Staying Engaged with Chapter 9 Chemical Names Formulas Answers Page 221
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 9 Chemical Names Formulas Answers Page 221
- 9. Balancing eBooks and Physical Books Chapter 9 Chemical Names Formulas Answers Page 221
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 9 Chemical Names Formulas Answers Page 221
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 9 Chemical Names Formulas Answers Page 221
  - Setting Reading Goals Chapter 9 Chemical Names Formulas Answers Page 221
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 9 Chemical Names Formulas Answers Page 221
  - Fact-Checking eBook Content of Chapter 9 Chemical Names Formulas Answers Page 221

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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