

## AP Chemistry Practice Test Chapter 8 and 9

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_ 1. Which pair of atoms should form the most polar bond?  
a. F and B  
b. C and O  
c. F and O  
d. N and F  
e. B and N
- \_\_\_\_ 2. Which pair of ions should form the ionic lattice with the highest energy?  
a.  $\text{Na}^{1+}$  and  $\text{Br}^{1-}$   
b.  $\text{Li}^{1+}$  and  $\text{F}^{1-}$   
c.  $\text{Cs}^{1+}$  and  $\text{F}^{1-}$   
d.  $\text{Li}^{1+}$  and  $\text{O}^{2-}$   
e.  $\text{K}^{1+}$  and  $\text{F}^{1-}$
- \_\_\_\_ 3. The electrostatic force of attraction is greatest in which compound?  
a.  $\text{BaO}$   
b.  $\text{MgO}$   
c.  $\text{CaS}$   
d.  $\text{MgS}$   
e.  $\text{CaO}$
- \_\_\_\_ 4. Which molecule has the weakest bond?  
a.  $\text{CO}$   
b.  $\text{O}_2$   
c.  $\text{NO}$   
d.  $\text{N}_2$   
e.  $\text{Cl}_2$
- \_\_\_\_ 5. For which species are octet resonance structures necessary and sufficient to describe the bonding satisfactorily?  
a.  $\text{BCl}_3$   
b.  $\text{SO}_2$   
c.  $\text{CO}_2$   
d.  $\text{BeF}_2$   
e.  $\text{SO}_4^{2-}$
- \_\_\_\_ 6. How are the bonding pairs arranged in the best Lewis structure for ozone?  
a.  $\text{O}-\text{O}-\text{O}$   
b.  $\text{O}=\text{O}-\text{O}$   
c.  $\text{O}=\text{O}=\text{O}$   
d.  $\text{O}=\text{O}=\text{O}$   
e.  $\text{O}=\text{O}=\text{O}$
- \_\_\_\_ 7. Which species has the shortest bond length?  
a.  $\text{CN}^{1-}$   
b.  $\text{O}_2$   
c.  $\text{SO}_2$   
d.  $\text{SO}_3$   
e.  $\text{CO}_2$
- \_\_\_\_ 8. Which species requires the least amount of energy to remove an electron from the outermost energy level?  
a.  $\text{Na}^{1+}$   
b. He  
c.  $\text{F}^{1-}$   
d.  $\text{O}^{2-}$   
e.  $\text{Mg}^{2+}$
- \_\_\_\_ 9. Which species has a valid non-octet Lewis structure?  
a.  $\text{GeCl}_4$   
b.  $\text{SiF}_4$   
c.  $\text{CO}_3^{2-}$   
d.  $\text{SeCl}_4$   
e.  $\text{CO}_3^{2-}$

# Chapter 9 Chemistry Test Answers

**Clifford Lane**



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