

50/50

Experiment 9 Prelaboratory Assignment

A Volumetric Analysis

Date 03/30

Desk No. _____

1. a. What is the titrant for this experiment?

NaOH (base), sodium hydroxide

b. Is the indicator generally added to the titrant or the analyte in a titration?

The indicator, phenolphthalein, is added to the analyte, the acid.
It is colorless in an acid but pink in basic solutions.

2. a. What is the primary standard used in this experiment (name and formula)? Define a primary standard.

A primary standard is an accurate mass of a solid substance which is measured, dissolved in water, and then reacted.

↳ $\text{KHC}_8\text{H}_4\text{O}_4$, potassium hydrogen phthalate

b. What is the secondary standard used in this experiment (name and formula)? Define a secondary standard.

A secondary standard is the standard remaining solution from after the molar concentration of the titrant is measured.

↳ NaOH, sodium hydroxide

3. Distinguish between a stoichiometric point and an endpoint in an acid-base titration.

reaction is complete
because amount
reactants that have
combined

when the indicator changes
color due to the change
in pH

4. a. When rinsing a buret after cleaning it with soap and water, should the rinse be dispensed through the buret tip or the top opening of the buret? Explain.

The water and rinse should go through the tip. Solution will dispense through and help keep the concentration of the titrant the same.



b. Experimental Procedure, Part A.4. In preparing the buret for titration, the final rinse is with the NaOH titrant rather than with deionized water. Explain.

The water can alter the values and the NaOH is hygroscopic.

5. Part A.3. Calculate the mass of $\text{KHC}_8\text{H}_4\text{O}_4$ (molar mass = 204.23 g/mol) that reacts with 15 mL of the 0.15 M NaOH solution prepared in Part A.2. Express this mass $\text{KHC}_8\text{H}_4\text{O}_4$ to the correct number of significant figures and record the calculation on the Report Sheet.

$$15 \text{ mL} = 0.015 \text{ L}$$

$$0.015 \text{ L (0.15 M)} = 0.00225 \text{ mol NaOH}$$

$$0.00225 \text{ mol NaOH} \cdot \frac{1 \text{ mol KHC}_8\text{H}_4\text{O}_4}{1 \text{ mol NaOH}} = 0.00225 \text{ mol KHC}_8\text{H}_4\text{O}_4$$

$$0.00225 \text{ mol KHC}_8\text{H}_4\text{O}_4 \cdot \left(\frac{204.23 \text{ g}}{1 \text{ mol}} \right) = \boxed{0.4595 \text{ g}}$$

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