

ChemActivity 40 B

Relationships among Equilibrium Constants

We saw in ChemActivity 40 that equilibrium constants can be calculated if similar reactions have known values of K .

1. Given that $2\text{NO}(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}_2(g)$; $K_p = 1.7 \times 10^{17}$ at 25°C

(a) Write the equilibrium expression $K_p =$

A similar reaction for which we can calculate K is: $\text{NO}(g) + \frac{1}{2}\text{O}_2(g) \rightleftharpoons \text{NO}_2(g)$.

(b) Write the equilibrium expression for the second reaction, $K_p' =$

(c) Describe the mathematical relationship between K_p and K_p' (how is K_p related to K_p' ?)

(d) What is the numerical value of K_p' ?

A third similar reaction is $\text{NO}_2(g) \rightleftharpoons \text{NO}(g) + \frac{1}{2}\text{O}_2(g)$

(e) Write the equilibrium expression for the third reaction, $K_p'' =$

(f) Describe the mathematical relationship between K_p' and K_p'' .

(g) What is the numerical value of K_p'' ?

In this course we have had need to add chemical reaction in a mathematical way. In our study of thermochemistry we added chemical reactions in applying Hess's Law. In each case all the reactants and products are combined and compounds that appear as both reactants and products are "canceled" from the final (overall) reaction. The same can be done with equilibrium expressions.

If we want to know the equilibrium constant for the atmospheric reaction



We can use reactions of known K that sum to give the desired equilibrium.

Chemactivity 40 Answers

Richard S. Moog, John J. Farrell



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Organic Chemistry, a Guided Inquiry Andrei Straumanis,2004 **Chemistry** Richard S. Moog,John J. Farrell,2017-06-26 In the newly updated 7th Edition Chemistry A Guided Inquiry continues to follow the underlying principles developed by years of extensive research on how students learn and draws on testing by those using the POGIL methodology This text follows the principles of inquiry based learning and correspondingly emphasizes underlying chemistry concepts and the reasoning behind them This text provides an approach that follows modern cognitive learning principles by having students learn how to create knowledge based on experimental data and how to test that knowledge

Introductory Chemistry Michael P. Garoutte,Ashley B. Mahoney,2015-08-10 The ChemActivities found in Introductory Chemistry A Guided Inquiry use the classroom guided inquiry approach and provide an excellent accompaniment to any one semester Introductory text Designed to support Process Oriented Guided Inquiry Learning POGIL these materials provide a variety of ways to promote a student focused active classroom that range from cooperative learning to active student participation in a more traditional setting **General, Organic, and Biological Chemistry** Michael P.

Garoutte,2014-02-24 Classroom activities to support a General Organic and Biological Chemistry text Students can follow a guided inquiry approach as they learn chemistry in the classroom General Organic and Biological Chemistry A Guided Inquiry serves as an accompaniment to a GOB Chemistry text It can suit the one or two semester course This supplemental text supports Process Oriented Guided Inquiry Learning POGIL which is a student focused group learning philosophy of instruction The materials offer ways to promote a student centered science classroom with activities The goal is for students to gain a greater understanding of chemistry through exploration **Merrill Chemistry** Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 Prentice Hall Chemistry ,2000

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