

## Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The charge on an ion is  
a. always positive.  
b. always negative.  
c. either positive or negative.  
d. zero.
- \_\_\_\_\_ 2. According to the octet rule, a calcium atom has a tendency to  
a. lose one electron.  
b. lose two electrons.  
c. gain one electron.  
d. gain two electrons.
- \_\_\_\_\_ 3. If a compound forms by ionic bonding, which is *not* true?  
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.  
b. The net charge of the compound is zero.  
c. The compound contains just two atoms, each of opposite charge.  
d. Several ions group together in a tightly packed structure.
- \_\_\_\_\_ 4. The only property listed that is *not* characteristic of ionic compounds is  
a. high melting point.  
b. hardness.  
c. lack of crystal structure.  
d. good conductor of electricity.
- \_\_\_\_\_ 5. Which formula listed below represents a polyatomic ion?  
a.  $\text{NH}_4^+$   
b.  $\text{H}_2\text{SO}_4$   
c.  $\text{NaCl}$   
d.  $\text{H}_2\text{O}$
- \_\_\_\_\_ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because  
a. ionic substances tend to vaporize at room temperature.  
b. ionic substances are brittle.  
c. attractive forces between ions are greater than the attractive forces between molecules.  
d. the numbers of positive and negative charges are equal in an ionic compound.

# Chemical Bonding Test With Answers

**RJ Shavelson**



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## **Table of Contents Chemical Bonding Test With Answers**

1. Understanding the eBook Chemical Bonding Test With Answers
  - The Rise of Digital Reading Chemical Bonding Test With Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Bonding Test With Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Platform
  - User-Friendly Interface
4. Exploring eBook Recommendations from eBook Platforms
  - Personalized Recommendations
  - eBook User Reviews and Ratings
  - eBook Bestseller Lists
5. Accessing eBook Chemical Bonding Test With Answers Free and Paid eBooks
  - eBook Public Domain eBooks
  - eBook Subscription Services

- Chemical Bonding Test With Answers Budget-Friendly Options
- 6. Navigating Chemical Bonding Test With Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Bonding Test With Answers Compatibility with Devices
  - Chemical Bonding Test With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Bonding Test With Answers
  - Highlighting and Note-Taking Chemical Bonding Test With Answers
  - Interactive Elements Chemical Bonding Test With Answers
- 8. Staying Engaged with Chemical Bonding Test With Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Bonding Test With Answers
- 9. Balancing eBooks and Physical Books Chemical Bonding Test With Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Bonding Test With Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Bonding Test With Answers
  - Setting Reading Goals Chemical Bonding Test With Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Bonding Test With Answers
  - Fact-Checking eBook Content of Chemical Bonding Test With Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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